

## GENERAL PRINCIPLES AND PROCESSES OF ISOLATION OF ELEMENTS

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- van-Arker method of purification of metals involves converting the metal to a
  - Volatile stable compound
  - Non-volatile stable compound
  - Volatile unstable compound
  - None of the above
- In the electrolysis of alumina, cryolite is added to:
  - Lower the melting point of alumina and to increase the electrical conductivity
  - Minimise the anode effect
  - Remove impurities from alumina
  - None of the above
- The pyrolusite ore contains:
  - Fe
  - Al
  - Mn
  - Cu
- Purest form of iron is
  - Pig iron
  - Wrought iron
  - Cast iron
  - Steel
- Pig iron is manufactured by:
  - An electric furnace
  - A blast furnace
  - An open hearth furnace
  - None of these
- During the process of electrolytic refining of copper, some metals present as impurity settle as 'anode mud'. These are
  - Fe and Ni
  - Ag and Au
  - Pb and Zn
  - Se and Ag
- By which process Pb and Sn are extracted respectively?
  - Carbon reduction—self reduction
  - Self reduction—carbon reduction
  - Electrolytic reduction—cyanide process
  - Cyanide process—electrolytic reduction
- CO on passing over heated nickel gives:
  - $\text{NiCO}_3$
  - $\text{Ni}(\text{CO})_4$
  - $\text{CO}_2 + \text{H}_2$
  - $\text{CO} + \text{H}_2$
- Cassiterite is concentrated by
  - Liquation
  - Floatation
  - Electromagnetic separation
  - Levigation
- In the extraction of copper from its sulphide ore, the metal is finally obtained by the reduction of cuprous oxide with:
  - Iron sulphide ( $\text{FeS}$ )
  - Carbon monoxide ( $\text{CO}$ )
  - Copper(I) sulphide ( $\text{Cu}_2\text{S}$ )
  - Sulphur dioxide ( $\text{SO}_2$ )
- Which of the following metal is thrown as anode mud during electrolytic refining of copper?
  - Zn
  - Fe
  - Ag
  - Ni
- Which metal is a liquid at room temperature?
  - Mercury
  - Potassium
  - Sodium
  - Titanium
- 'Lapis-Lazuli' is a blue coloured precious stone. It is mineral of the class
  - Sodium alumino silicate
  - Basic copper carbonate





33. The mass of carbon anode consumed (giving only carbon dioxide) in the production of 270 kg of aluminium metal from bauxite by the hall process is  
(Atomic mass of Al=27)
- a) 180kg                      b) 270 kg                      c) 540 kg                      d) 90 kg
34. Carbon monoxide reduction process is used for the extraction of:
- a) Cu                      b) Ag                      c) Na                      d) K
35. Lead stone is one ore of
- a) Iron                      b) Lead                      c) Silicon                      d) Tin
36. One of the following metals forms a volatile compound and this property is taken advantage for its extraction. This metals is
- a) Cobalt                      b) Iron                      c) Tungsten                      d) Nickel
37. Carbon reduction is used for the extraction of:
- a) Fe                      b) K                      c) Al                      d) None of these
38. The phenomenon in which white transparent crystal changes into white powder is known as:
- a) Sublimation                      b) Allotropy                      c) Efflorescence                      d) deliquescence
39. Which is used for the extraction of cadmium from cadmium sulphide?
- a) Roasting                      b) Reduction                      c) Oxidation                      d) Electrolysis
40. Formula of magnetite is
- a)  $\text{Fe}_3\text{O}_4$                       b)  $\text{Fe}_2\text{O}_3$                       c)  $\text{FeS}_2$                       d)  $\text{FeCO}_3$
41. When  $\text{MnO}_2$  is fused with KOH, a coloured compound is formed, the compound and its colour is:
- a)  $\text{K}_2\text{MnO}_4$ , purple green  
b)  $\text{KMnO}_4$ , purple  
c)  $\text{Mn}_2\text{O}_3$ , brown  
d)  $\text{Mn}_3\text{O}_4$ , black
42. Which is not a basic flux?
- a)  $\text{CaCO}_3$                       b) CaO                      c)  $\text{SiO}_2$                       d) MgO
43. An ore of tin containing  $\text{FeCrO}_4$  is concentrated by:
- a) Magnetic separation                      b) Froth floatation                      c) Electrostatic method                      d) Gravity separation
44. Orford process is used in extraction of:
- a) Pt                      b) Co                      c) Fe                      d) Ni
45. The salt which is least likely to be found in minerals is:
- a) Sulphate                      b) Acetate                      c) Chloride                      d) Sulphide
46. The second most common element on the earth is:
- a) Silicon                      b) Hydrogen                      c) Nitrogen                      d) Oxygen
47. An ore of tin containing  $\text{FeCrO}_4$  is concentrated by
- a) Electrostatic method                      b) Gravity separation                      c) Magnetic separation                      d) Forth floatation
48. Alkaline earth metals are not found free in nature because of:
- a) Their high b. p.  
b) Their low b. p.  
c) Thermal instability  
d) Their great chemical activity
49. Alloy is an example of:
- a) Gel                      b) Aerosol                      c) Solid sol                      d) Emulsion
50. Cinnabar is an ore of
- a) Pb                      b) Hg                      c) Cu                      d) Zn
51. Which element occurs in free state in nature?
- a) Fe                      b) Co                      c) Pt                      d) Ni
52. Aluminothermic process is used for the extraction of metals, whose oxides are:
- a) Fusible  
b) Not easily reduced by carbon

- c) Not easily reduced by hydrogen  
d) Strongly basic
53. Bauxite ore is concentrated by  
a) Froth floatation  
b) Electromagnetic separation  
c) Chemical separation  
d) Hydraulic separation
54. Which process is used for beneficiation of ores?  
a) Process of removal of impurities  
b) Process of heating ore at high temperature  
c) Extraction of metal from ore  
d) None of the above
55. Extraction for zinc from zinc blende is achieved by  
a) Electrolytic reduction  
b) Roasting followed by reduction with carbon  
c) Roasting followed by reduction with another metal  
d) Roasting followed by self-reduction
56. Auto-reduction process is used in the extraction of  
a) Cu and Hg  
b) Zn and Hg  
c) Cu and Al  
d) Fe and Pb
57. Thomas slag is  
a)  $\text{Ca}_3(\text{PO}_4)_2 \cdot 2\text{H}_2\text{O}$   
b)  $\text{Ca}_3(\text{PO}_4)_2 \cdot \text{CaSiO}_3$   
c)  $\text{MgSiO}_3$   
d)  $\text{CaSiO}_3$
58. Metals are good conductors of electricity because they contain  
a) Ionic bonds  
b) A network structure  
c) Very few valence electrons  
d) Free electrons
59. Liquation is used to purify:  
a) Hg  
b) Sn  
c) Bi  
d) All of these
60. The most abundant metal in the earth crust is:  
a) Na  
b) Ca  
c) Al  
d) Fe
61. Which of the elements listed below shows allotropic forms?  
a) Iodine  
b) Copper  
c) Sulphur  
d) Silver
62. Following method is not used for extraction of Al  
a) Van Arkel  
b) Serpeck  
c) Baeyer  
d) Hall-Heroult
63. Indian saltpetre is:  
a)  $\text{KNO}_2$   
b)  $\text{KNO}_3$   
c) NaCl  
d)  $\text{Na}_2\text{CO}_3$
64. Poling process is used:  
a) For the removal of  $\text{Cu}_2\text{O}$  from Cu  
b) For the removal of  $\text{Al}_2\text{O}_3$  from Al  
c) For the removal of  $\text{Fe}_2\text{O}_3$  from Fe  
d) In all of the above
65. Sperrylite is:  
a) AgCl  
b)  $\text{PtAs}_2$   
c)  $\text{Fe}_2\text{O}_3$   
d)  $\text{SnO}_2$
66. The substance added in water in the froth floatation process is  
a) Pine oil  
b) Coconut oil  
c) Soap powder  
d) None of these
67. The region in which metals are found in earth is called:  
a) Atomophil  
b) Lithophil  
c) Calcophil  
d) Sidrophil
68. In the manufacture of iron from haematite, lime stone is added to act as:  
a) Flux  
b) Slag  
c) A reducing agent  
d) An oxidising agent
69. On heating a mixture of  $\text{Cu}_2\text{O}$  and  $\text{Cu}_2\text{S}$ , we get:  
a)  $\text{Cu} + \text{SO}_2$   
b)  $\text{Cu} + \text{SO}_3$   
c)  $\text{CuO} + \text{CuS}$   
d)  $\text{Cu}_2\text{SO}_3$
70. Cassiterite is an ore of  
a) Sb  
b) Mn  
c) Sn  
d) Ni
71. In the metallurgical extraction of zinc from  $\text{ZnO}$ , the reducing agent used is?



- a) Nitric oxide                      b) Sulphur dioxide                      c) Carbon monoxide                      d) Carbon dioxide
72. Zinc blende (an ore) is:  
a) ZnO                      b) ZnCO<sub>3</sub>                      c) ZnS                      d) Zn<sub>2</sub>OCl<sub>2</sub>
73. From gold amalgam, gold may be recovered by:  
a) Addition of Zn metal  
b) Electrolytic refining  
c) Distillation  
d) Dissolving Hg in HNO<sub>3</sub>
74. The lightest metal is:  
a) Li                      b) Mg                      c) Ca                      d) Na
75. Calamine is  
a) CaCO<sub>3</sub>                      b) MgCO<sub>3</sub>  
c) ZnCO<sub>3</sub>                      d) CaCO<sub>3</sub> + CaO
76. In the metallurgy of iron, when lime stone is added to the blast furnace, the calcium ions are removed as:  
a) Slag                      b) Gangue                      c) Metallic Ca                      d) CaCO<sub>3</sub>
77. Mond's process is used for the purification of  
a) Ni                      b) Ti                      c) Zr                      d) Hg
78. Which contains both Ca and Mg?  
a) Lime stone                      b) Dolomite                      c) Chalk                      d) Felspar
79. Calcination and roasting are:  
a) Different names of the same operation  
b) Used for the purification of metals  
c) Usually carried out in reverberatory furnace  
d) Employed for the concentration of the ore
80. Passivity of iron is due to the formation of thin film of....on its surface.  
a) Oxide                      b) Carbonate                      c) Nitride                      d) Hydroxide
81. Which of the following statement is incorrect?  
a) Silver glance mainly contains silver sulphide                      b) Zinc blende mainly contains zinc chloride  
c) Gold is found in native state                      d) Copper pyrites also contains Fe<sub>2</sub>S<sub>3</sub>
82. Of the following, which cannot be obtained by electrolysis of the aqueous solution of their salts?  
a) Cu                      b) Ag                      c) Mg and Al                      d) Cr
83. The sand stone in some iron ores is removed by:  
a) Carbon filters                      b) Compressed air                      c) Lime stone                      d) Sulphuric acid
84. Copper pyrites is concentrated by  
a) Gravity method                      b) Forth floatation process  
c) Electromagnetic method                      d) All of these
85. The chief impurity present in red bauxite is  
a) SiO<sub>2</sub>                      b) Fe<sub>2</sub>O<sub>3</sub>                      c) K<sub>2</sub>SO<sub>4</sub>                      d) NaF
86. Which does not contain aluminium?  
a) Bauxite                      b) Emery                      c) Rutile                      d) Corundum
87. Naturally occurring substances from which a metal can be profitably (or economically) extracted are called  
a) Ores                      b) Mineral                      c) Salts                      d) Gangue
88. Ferric oxide in blast furnace is reduced by:  
a) C                      b) H<sub>2</sub>                      c) CO                      d) CO<sub>2</sub>
89. Cupellation process is used in the metallurgy of:  
a) Copper                      b) Silver                      c) Lead                      d) Iron
90. Which metal can be purified by distillation?  
a) Cu                      b) Ag                      c) Fe                      d) Hg
91. Lepidolite is an ore of:



- a) K                                      b) Na                                      c) Li                                      d) All of these
92. Chalcogens are:  
 a) Hydrocarbons  
 b) Ore forming elements  
 c) Oxide forming elements  
 d) Those having ability to catenate
93. In the Hall's process for extraction of Al, the ore is fused with:  
 a)  $\text{NaHCO}_3$                                       b)  $\text{Na}_2\text{CO}_3$                                       c) NaF                                      d)  $\text{Na}_3\text{AlF}_6$
94. Antimony occurs mainly in form of:  
 a) Sulphide                                      b) Stibnite                                      c) Realgar                                      d) Fluorapatite
95. An important ore of iron is  
 a) Pyrites                                      b) Malachite                                      c) haematite                                      d) Siderite
96. Barytes, an ore is:  
 a)  $\text{BeSO}_4$                                       b)  $\text{BeCl}_2$                                       c)  $\text{BaSO}_4$                                       d)  $\text{BaCl}_2$
97. Thermite is a mixture of  
 a) Fe powder and  $\text{Al}_2\text{O}_3$                                       b) Al powder and  $\text{Fe}_2\text{O}_3$   
 c) Cu powder and  $\text{Fe}_2\text{O}_3$                                       d) Zn powder and  $\text{Cr}_2\text{O}_3$
98. When lime stone is heated, carbon dioxide is given off. This operation in metallurgy is known as:  
 a) Smelting                                      b) Ore-dressing                                      c) Calcination                                      d) Roasting
99. Heating mixture of  $\text{Cu}_2\text{O}$  and  $\text{Cu}_2\text{S}$  will give  
 a)  $\text{Cu}_2\text{SO}_3$                                       b)  $\text{CuO} + \text{CuS}$                                       c)  $\text{Cu} + \text{SO}_3$                                       d)  $\text{Cu} + \text{SO}_2$
100. A and B are two allotropes of an element. One gram of A will differ from one gram of B in:  
 a) Oxidation number  
 b) Chemical composition  
 c) Total number of atoms  
 d) Atomic arrangement
101. Which represents calcination?  
 a)  $2\text{Ag} + 2\text{HCl} + [\text{O}] \rightarrow 2\text{AgCl} + \text{H}_2\text{O}$   
 b)  $2\text{Zn} + \text{O}_2 \rightarrow 2\text{ZnO}$   
 c)  $2\text{ZnS} + 3\text{O}_2 \rightarrow 2\text{ZnO} + 2\text{SO}_2$   
 d)  $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$
102. The matte is impure substance obtained during extraction of:  
 a) Cu                                      b) Fe                                      c) Pb                                      d) Al
103. The following equation represents a method of purification of nickel by,  

$$\text{Ni} + 2\text{CO} \xrightarrow{320\text{K}} \text{Ni}(\text{CO})_4 \xrightarrow{420\text{K}} \text{Ni} + 4\text{CO}$$
 Impure                                      Pure  
 This method is:  
 a) Cupellation                                      b) Mond's process                                      c) Van Arkel method                                      d) Zone refining
104. Softening of lead means:  
 a) Conversion of lead into PbO  
 b) Conversion of lead into  $\text{Pb}_3\text{O}_4$   
 c) Removal of metallic impurities from lead  
 d) Washing lead with  $\text{HNO}_3$  followed by dilute alkali solution
105. Which is not a mineral of aluminium?  
 a) Corundum                                      b) Anhydrite                                      c) Diaspore                                      d) Bauxite
106. A common metal used as reductant for the extraction of metals from their oxides is:  
 a) Cr                                      b) Al                                      c) Co                                      d) Fe
107. Extraction of Ag from commercial lead is possible by:  
 a) Parke's process                                      b) Clarke's process                                      c) Pattinson's process                                      d) Electrolytic process
108. Which set of elements is called chalcogens?



- a) Cl, Br, I                      b) O, S, Se                      c) N, P, As                      d) C, Si, Ge
109. Apatite is an ore of  
a) Fluorine                      b) Chlorine                      c) Bromine                      d) iodine
110. Pentlandite is an ore of:  
a) Fe                      b) Co                      c) Cu                      d) Ni
111. Which element is present in pitch blende?  
a) U                      b) Ce                      c) Ba                      d) Mg
112. In aluminothermite process, aluminium is used as  
a) Reducing agent                      b) Oxidizing agent                      c) Solder                      d) Flux
113. The existence of two or more crystalline forms of the same substance is called:  
a) Polymorphism                      b) Isomerism                      c) Homologues                      d) Isomorphism
114. Froth floatation process for the concentration of the ores is an illustration of the practical application of  
a) Adsorption                      b) Sedimentation                      c) Coagulation                      d) Absorption
115. In blast furnace, the cup and cone arrangement is used:  
a) To escape the gases during charging  
b) Not to allow the escape of the gases  
c) To heat the charge with the gases  
d) None of the above
116. Stainless steel has iron and  
a) Cr                      b) Cu                      c) Co                      d) Zn
117. Blood haemoglobin contains:  
a) Al                      b) Mg                      c) Cu                      d) Fe
118. Cyanide process is used in the extraction of  
a) Au                      b) Cu                      c) Ag                      d) Both (a) and (c)
119. Alloy formation gives rise to:  
a) Decrease in corrosion  
b) Increase in hardness  
c) Decrease in conductivity  
d) All are correct
120. Which metal occurs in free state?  
a) Ag                      b) Au                      c) Pt                      d) All of these
121. Platinum, palladium, indium, etc., are called noble metals because:  
a) Alfred nobel discovered them  
b) They are inert towards many common reagents  
c) They are shining, lustrous and pleasing to look at  
d) They are found in native state
122. Match the extraction process listed in column I with metals listed in column II.
- | Column I                                       | Column II  |
|--|------------|
| A. Self reduction                              | (P) Lead   |
| B. Carbon reduction                            | (Q) Silver |
| C. Complex formation and displacement by metal | (R) Copper |
| D. Decomposition of iodide                     | (S) Boron  |
- a) A – P, R; B – R, Q; C – P; D – S, Q  
b) A – P, R; B – P, R; C – Q; D – S  
c) A – P, R; B – S; C – P; D – P, Q  
d) A – P, Q; B – R, P; C – Q; D – S
123. Mercury is transported in metal containers made up of:  
a) Fe                      b) Pb                      c) Zn                      d) Sn
124. Which is not a mineral?



- a) Mica                      b) Peat                      c) Quartz                      d) Felspar
125. Slag coming out at the bottom of a blast furnace during extraction of iron from its ores, is used in making:  
a) Roads                      b) Fertilizers                      c) Plastics                      d) Glass moulds
126. The process in which ore is heated in air below its melting point is known as:  
a) Roasting                      b) Calcination                      c) Reduction                      d) Distillation
127. When pyrolusite is fused with KOH in presence of air, the fused mass becomes:  
a) Pink                      b) Green                      c) Red                      d) Black
128. Which process is used for the purification of Al metal?  
a) Hoop's process                      b) Baeyer's process                      c) Serpek's process                      d) Hall's process
129. Which is incorrect as the uses of lime stone in industries are concerned?  
a) For making cement  
b) In the extraction of Sn from its ore  
c) In the extraction of Fe from its ore  
d) In the manufacture of glass
130. The method of zone refining of metals is based on the principle of  
a) Greater noble character of the solid metal than that of the impurity  
b) Greater solubility of the impurity in the molten state than in the solid  
c) Greater mobility of the pure metal than that of impurity  
d) Higher melting point of the impurity than that of the pure metal
131. Main ore of aluminium is:  
a) Cryolite                      b) Kaolin                      c) Bauxite                      d) Felspar
132. Which of the following is a carbonate ore?  
a) Pyrolusite                      b) Diaspore                      c) Cassiterite                      d) Malachite
133. Which of the following mineral does not contain Al?  
a) Fluorspar                      b) Cryolite                      c) Mica                      d) Feldspar
134. An essential constituent of amalgam is:  
a) Au                      b) Ag                      c) Al                      d) Hg
135. Mispickel is the ore of:  
a) Sb                      b) Bi                      c) P                      d) As
136. Froth floatation method is successful in separating impurities from ores because  
a) The pure ore is soluble in water containing additives like pine oil, cresylic acid etc  
b) The pure ore is lighter than water containing additives like pine oil, cresylic acid, etc  
c) The impurities are soluble in water containing additives like pine oil, cresylic acid, etc  
d) The pure ore is not easily wetted by water as by pine oil, cresylic acid, etc
137. Which among the following has highest electrical conductivity?  
a) Zn                      b) Fe                      c) Ag                      d) Cu
138. Which of the following statements regarding the metallurgy of magnesium using electrolytic method is not correct?  
a) Electrolyte is magnesium chloride containing a little of NaCl and NaF  
b) Air tight iron pot acts as a cathode  
c) Electrolysis is done in the atmosphere of coal gas  
d) Molten magnesium is heavier than the electrolyte
139. The process of heating the ore strongly in excess of air so that the volatile impurities are removed and the ore is changed to oxide is known as  
a) Leaching                      b) Roasting                      c) Calcinations                      d) Froth floatation
140. During bessemerisation of copper, the reaction taking place in the bessemer convertor is:  
a)  $\text{Cu}_2\text{S} + 2\text{Cu}_2\text{O} \rightarrow 6\text{Cu} + \text{SO}_2$   
b)  $\text{Cu}_2\text{O} + \text{FeS} \rightarrow \text{Cu}_2\text{S} + \text{FeO}$   
c)  $\text{FeO} + \text{SiO}_2 \rightarrow \text{FeSiO}_3$   
d) None of the above



141. Which process is used for the extraction of metals from their sulphide ores?  
 a) Electrolysis                      b) Metal displacement                      c) Smelting                      d) Roasting
142. When copper pyrites is roasted in excess of air, a mixture of  $\text{CuO} + \text{FeO}$  is formed.  $\text{FeO}$  is present as impurities. This can be removed as slag during reduction of  $\text{CuO}$ . The flux added to form slag is  
 a)  $\text{SiO}_2$  which is an acid flux                      b) Lime stone, which is a basic flux  
 c)  $\text{SiO}_2$ , which is basic flux                      d)  $\text{CaO}$ , which is basic flux
143.  $\text{CaO}$  act as ... flux  
 a) Neutral                      b) Acidic                      c) Basic                      d) Both (a) and (b)
144. Electrolysis of fused carnallite gives:  
 a) Mg                      b) K                      c) K and  $\text{CO}_2$                       d) K, Mg and  $\text{Cl}_2$
145. Wolframite ore is separated from tin stone ore by the process of  
 a) Calcination                      b) Electromagnetic                      c) Roasting                      d) Smelting
146. Iron ores are dressed by:  
 a) Froth floatation process  
 b) Magnetic separation  
 c) Hand picking  
 d) All of the above
147. The electrolytic reduction technique is used in the extraction of:  
 a) Highly electronegative elements  
 b) Highly electropositive elements  
 c) Metalloids  
 d) Transition metals
148. Iron is obtained on large scale from  $\text{Fe}_2\text{O}_3$  by:  
 a) Reduction with  $\text{CO}$                       b) Reduction with Al                      c) Calcination                      d) Passing  $\text{H}_2$
149. The lining in blast furnace are made up of:  
 a) Graphite                      b) Silica                      c) Fireclay bricks                      d)  $\text{CaCO}_3$
150. The cyanide process is used for obtaining  
 a) Cu                      b) Na                      c) Zn                      d) Ag
151. Refractory materials are used for the construction of furnaces because they:  
 a) Are light in weight  
 b) Can stand with high temperature  
 c) Are leak proof  
 d) Do not require to be replaced
152. The final step for the extraction of copper from copper pyrite in Bessemer converter involves the reaction  
 a)  $\text{Cu}_2\text{S} + 2\text{Cu}_2\text{O} \rightarrow 6\text{Cu} + \text{SO}$                       b)  $4\text{Cu}_2\text{O} + \text{FeS} \rightarrow 8\text{Cu} + \text{FeSO}_4$   
 c)  $2\text{Cu}_2\text{O} + \text{FeS} \rightarrow 4\text{Cu} + \text{Fe} + \text{SO}_2$                       d)  $\text{Cu}_2\text{S} + 2\text{FeO} \rightarrow 2\text{Cu} + 2\text{FeCO} + \text{SO}_2$
153. Beryl is an important ore of:  
 a) Boron                      b) Beryllium                      c) Lead                      d) Lithium
154. Smelting is done in:  
 a) Blast furnace                      b) Muffle furnace                      c) Open hearth furnace                      d) Electric furnace
155. Silver obtained by argentiferrous lead is purified by:  
 a) Distillation                      b) Froth floatation                      c) Cupellation                      d) Reacting with KCN
156. Among the following groups of oxides, the group containing oxides that cannot be reduced by carbon to give the respective metals is  
 a)  $\text{Cu}_2\text{O}, \text{K}_2\text{O}$                       b)  $\text{PbO}, \text{Fe}_3\text{O}_4$                       c)  $\text{Fe}_2\text{O}_3, \text{ZnO}$                       d)  $\text{CaO}, \text{K}_2\text{O}$
157. Which metal can be found in native state?  
 a) Na                      b) Al                      c) Ca                      d) Fe
158. Which of the following pairs of metals is purified by van Arkel method?  
 a) Ni and Fe                      b) Ga and In                      c) Zr and Ti                      d) Ag and Au
159. Which of the following is the heaviest metal?

- a) U                                      b) Ra                                      c) Pb                                      d) Hg
160. Iron is made inactive or passive by:  
a)  $H_3PO_4$                                       b) Conc.  $HNO_3$                                       c) Conc.  $H_2SO_4$                                       d) Dil.  $HNO_3$
161. Kiesserite is an ore of:  
a) Cu                                      b) Al                                      c) Mg                                      d) Fe
162. Smelting is the reduction of oxide to metal by:  
a) C                                      b) Al                                      c) H                                      d) Electric current
163. Which of the following is a metal?  
a) P                                      b) As                                      c) Bi                                      d) Sb
164. Sulphide ores of metals are usually concentrated by froth floatation process. Which one of the following sulphide ores offers an exception and is concentrated by leaching?  
a) Galena                                      b) Copper pyrite                                      c) Sphalerite                                      d) Argentite
165. Which consists of only one element?  
a) Marble                                      b) Sand                                      c) Diamond                                      d) Glass
166. Impurities physically associated with minerals are:  
a) Slag                                      b) Flux                                      c) Alloy                                      d) Matrix
167. One of the fertilizer is:  
a)  $CaC_2$                                       b)  $CaCO_3$                                       c)  $CaCN_2$                                       d)  $CaSO_4$
168. In the commercial electrochemical process for aluminium extraction, electrolyte used is:  
a)  $Al(OH)_3$  in NaOH solution  
b) An aqueous solution of  $Al_2(SO_4)_3$   
c) A molten mixture of  $Al_2O_3$  and  $Na_3AlF_6$   
d) A molten mixture of  $Al_2O_3$  and  $Al(OH)_3$
169. Which element is found in human body?  
a) Pb                                      b) Fe                                      c) Cd                                      d) Al
170. Flux is used to remove  
a) Acidic impurities                                      b) Basic impurities  
c) All impurities from ores                                      d) From ores
171. Which statement is correct?  
a) Slag are carefully chosen to combine with the slag present in the ore to produce easily fusible gangue to carry away the impurities  
b) Gangues are carefully chosen to combine with the slag present in the ore to produce easily fusible flux to carry away the impurities  
c) Gangues are carefully chosen to combine with flux present in the ore to produce easily fusible slag to carry away the impurities  
d) Fluxes are carefully chosen to combine with the gangue present in the ore to produce easily fusible slag to carry away the impurities
172. Thermite process is used in reduction of  
a)  $Cr_2O_3$                                       b)  $Al_2O_3$                                       c)  $PbO_2$                                       d) CuO
173. Froth floatation process for the concentration of ores is a practical application of:  
a) Adsorption                                      b) Absorption                                      c) Coagulation                                      d) Sedimentation
174. The main constituent of steel in India are:  
a) Ni and Mg                                      b) V and Co                                      c) Al and Zn                                      d) Mn and Cr
175. Which is not employed for refining of metal?  
a) Poling                                      b) Leaching                                      c) Electrolysis                                      d) Liquation
176. In electrofining of copper, some gold is deposited as  
a) Cathode                                      b) Electrode                                      c) Cathode mud                                      d) Anode mud
177. Electric furnaces are lined with magnesia because:  
a) It is not affected by acids  
b) It liberates oxygen on heating



- c) It melts at very high temperature  
d) It has no effect of electricity
178. When the sample of Cu with Zn impurity is to be purified by electrolysis, the appropriate electrodes are:
- | Cathode          | Anode         |
|------------------|---------------|
| a) Pure Zn       | Pure Cu       |
| b) Impure sample | Pure Cu       |
| c) Impure Zn     | Impure sample |
| d) Pure Cu       | Impure sample |
179. The process of concentrating silver ore is based on its solubility in:
- a) HCl                                      b) HNO<sub>3</sub>                                      c) KCN                                      d) NaOH
180. Correct statement is
- a) van-Arkel method is used for extraction of Zr                      b) Limestone is acidic flux  
c) Dolomite is an ore of Al                                      d) Willemite is carbonate ore
181. Which one of the following ores is best concentrated by forth-floatation method?
- a) Magnetite                                      b) Cassiterite                                      c) Galena                                      d) Malachite
182. Boron is found in form of:
- a) Borax                                      b) Colemanite                                      c) Both (a) and (b)                      d) None of these
183. Extraction of silver from its ore involving NaCN, air and an active metal is known as:
- a) Pattinson's method  
b) Amalgamation method  
c) Mc Arthur-Forest method  
d) Parke's method
184. Heating of ore in presence of air to remove impurity of sulphur is called:
- a) Calcination                                      b) Roasting                                      c) Smelting                                      d) None of these
185. The ore concentrated by electromagnetic separation is:
- a) Wolframite                                      b) Haematite                                      c) Casseterite                                      d) All of these
186. Which process represents the change,  
 $Ti + 2I_2 \rightarrow TiI_4 \rightarrow Ti + 2I_2$  ?
- a) Cupellation                                      b) Van Arkel                                      c) Poling                                      d) Zone refining
187. Liquid crystals are best used in:
- a) Colour TV                                      b) Crystallization                                      c) Extraction                                      d)  $e/m$  determination
188. In the metallurgy of zinc, the zinc dust obtained from roasting and reduction of zinc sulphide contains some ZnO. It is removed by:
- a) Absorbance of ultraviolet light and reemission of white light  
b) Shock cooling by contact with a shower of molten lead  
c) X-ray method  
d) Smelting
189. High purity copper metal is obtained by:
- a) Carbon reduction                      b) Hydrogen reduction                      c) Electrolytic reduction                      d) Thermite process
190. In metallurgy, flux is a substance used to convert
- a) Soluble impurities to insoluble impurities                      b) Infusible impurities to fusible material  
c) Fusible impurities to infusible impurities                      d) Mineral into silicate
191. Gold is found usually near....mineral.
- a) Mica                                      b) Felspar                                      c) Quartz                                      d) Galena
192. The smelting of iron in a blast furnace involves all the steps except:
- a) Reduction                                      b) Fusion                                      c) Decomposition                                      d) Sublimation
193. The metal that is extracted from sea water is:
- a) Na                                      b) Ca                                      c) Mg                                      d) Sn
194. Wulfenite (a yellow-red mineral) having waxy lustre occur in lead ores, is an important source of:
- a) Sulphur                                      b) Molybdenum                                      c) Helium                                      d) Lead

195. The froth-floatation process is based upon
- The difference in the specific gravity of ore and gangue particles
  - The magnetic properties of gangue and ore
  - Preferential wetting of gangue particles by oil
  - The solubility of ore particles in suitable reagent
196. Pig iron is converted into steel by decreasing the amount of carbon contained in it, in a:
- Blast furnace
  - Pyrite burner
  - Bessemer converter
  - None of these
197. Plumbo-solvency refers to:
- Oxidation of lead to lead oxide
  - Oxidation of lead to red lead
  - Dissolution of lead in water containing air
  - Making lead wires by forcing heated metal through a die
198. Zinc is obtained on large scale by:
- Electrolysis of  $ZnCl_2$
  - Reduction of  $ZnO$
  - Precipitation with Ag
  - All are correct
199. Which of the following substances can be used for drying gases?
- $CaO$
  - $NaHCO_3$
  - $CaCO_3$
  - $Na_2CO_3$
200. Refractory materials are generally used in furnaces because
- They can withstand high temperature
  - They are chemically inert
  - They do not require replacement
  - They possess great structural strength
201. Presence of small impurity usually makes a metal quite hard because the impurities:
- Change the lattice structure of metals
  - Reduce the number of slide planes
  - Reduce the number of mobile electrons
  - Reduce the crystal symmetry
202. Willemite is
- $Zn_2SiO_4$
  - $H_2PtCl_6$
  - $ZnO$
  - $ZnOFe_2O_3$
203. The least stable oxide at room temperature is:
- $ZnO$
  - $CuO$
  - $Sb_2O_3$
  - $Ag_2O$
204. The process of removal of gangue particles from ores is known as:
- Concentration
  - Refining
  - Smelting
  - None of these
205. The process of calcination and roasting are carried out in:
- Blast furnace
  - Muffle furnace
  - Reverberatory furnace
  - Open hearth furnace
206. Which is not essential for rusting?
- Oxygen
  - Water
  - Carbon dioxide
  - Iron
207. Which of the following does not contain silicon?
- Kaoline
  - Agate
  - Ruby
  - Quartz
208. The salt which is least likely to be found in minerals is:
- Chloride
  - Sulphate
  - Sulphide
  - Nitrate
209. Heating of pyrite ores in air to remove Sulphur is known as:
- Calcination
  - Fluxing
  - Smelting
  - Roasting
210. Leaching is a process of:
- Reduction
  - Concentration
  - Refining
  - Oxidation
211. Colemanite is
- $Ca[B_3O_4(OH)_2] \cdot 2H_2O$
  - $Ca_2B_6O_{11} \cdot 5H_2O$
  - $Ca(OH)_2$
  - $Na_2B_4O_7 \cdot 2H_2O$
212. The ore that is concentrated by froth floatation process is
- Zincite
  - Cinnabar
  - Bauxite
  - malachite



213. Which one of the following ores is a chloride?  
 a) Bauxite                      b) Horn silver                      c) Zincite                      d) Felspar
214. An example of an oxide is  
 a) Zinc blende                      b) Bauxite                      c) Feldspar                      d) Malachite
215. The chemical composition of carnallite is:  
 a)  $\text{KCl} \cdot \text{MgCl}_2 \cdot 6\text{H}_2\text{O}$                       b)  $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$                       c)  $\text{MgCO}_3 \cdot 7\text{H}_2\text{O}$                       d)  $\text{MgCO}_3$
216. Which is not a silver ore?  
 a) Argentite                      b) Siderite                      c) Horn silver                      d) Ruby silver
217. Blast furnace is used in the metallurgy of:  
 a) Al                      b) Fe                      c) Gold                      d) Ag
218. Corundum is  
 a)  $\text{Cu}_2\text{Cl}_2$                       b)  $\text{CaCl}_2$                       c)  $\text{SrO}_2$                       d)  $\text{Al}_2\text{O}_3$
219. An alloy is:  
 a) Intermetallic compound  
 b) A solid substance containing two or more elements  
 c) A solid which contains one non-metal  
 d) A solid which contains more than one non-metal
220. Which of the following is not ore?  
 a) Zinc blende                      b) Malachite                      c) Bauxite                      d) Pig iron
221. Cryolite is  
 a) Sodium borofluoride                      b) Magnesium silicate  
 c) Aluminium                      d) Sodium aluminium fluoride
222. In the thermite process the reducing agent is:  
 a) C                      b) Al                      c) Na                      d) Mg
223. Which is not an ore of lead?  
 a) Galena                      b) Cassiterite                      c) Anglesite                      d) Cerussite
224. Which is not an ore of nickel?  
 a) Nickel glance                      b) Garnerite                      c) Haematite                      d) Pentlandite
225. The ore magnesite is:  
 a)  $\text{MgCO}_3 \cdot \text{CaCO}_3$                       b)  $\text{MgCl}_2 \cdot \text{KCl} \cdot 6\text{H}_2\text{O}$                       c)  $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$                       d)  $\text{MgCO}_3$
226. In blast furnace, the highest temperature is in  
 a) Fusion zone                      b) Reduction zone                      c) Combustion zone                      d) Slag zone
227. Which one of the following is correct?  
 a) All minerals are ores                      b) All ores cannot be a mineral  
 c) A mineral cannot be an ore                      d) All ores are minerals
228. Furnaces are lined with calcium oxide because:  
 a) It gives off oxygen on heating  
 b) It gives light on heating  
 c) It is refractory and basic  
 d) It is not affected by acids
229. Lepidolite, a lithium ore, also contains:  
 a) Ru                      b)  $\text{MgSO}_4$                       c) Na                      d) Cs
230. Gold when dissolved in aqua-regia gives:  
 a) Auric chloride                      b) Aurous chloride                      c) Chloroauric acid                      d) Tempering
231. Specific gravity of slag is:  
 a) Always higher than molten metal  
 b) Always less than molten metal  
 c) Same as that of molten metal  
 d) None of the above
232. The correct statement is:



- a) Dolomite is the ore of zinc  
 b) Galena is the ore of mercury  
 c) Pyrolusite is the ore of iron  
 d) Cassiterite is the ore of tin
233. Which is known as blister copper?  
 a) Pure copper                      b) 98% copper                      c) Ore of copper                      d) Alloy of copper
234. Which of the following ore is not concentrated by froth floatation process?  
 a) Pyrolusite                      b) Pentlandite                      c) Zinc blende                      d) Copper pyrites
235. The metal extracted by leaching with cyanide is:  
 a) Mg                      b) Ag                      c) Cu                      d) Na
236. Dolomite is an ore of:  
 a) Li                      b) Rb                      c) K                      d) Cs
237. Which statement is incorrect?  
 a) Galena is an ore of Pb  
 b) Electrostatic separation is used for lead sulphide  
 c) Ore is heated strongly, above its melting point in roasting  
 d) Silica acts as acidic flux
238. Anglesite is an ore of:  
 a) Cd                      b) Ni                      c) Sb                      d) Pb
239. Froth floatation process is based on:  
 a) Wetting properties of ore particles  
 b) Specific gravity of ore particles  
 c) Magnetic properties of ore particles  
 d) Electrical properties of ore particles
240. In froth floatation process many chemicals (frother, collector, activator and depressant) are used. Which is called a frother?  
 a)  $\text{CuSO}_4$                       b)  $\text{NaCN} + \text{alkali}$                       c) Pine oil                      d) Potassium xanthate
241. Which metal is used as a reducing agent in smelting?  
 a) C                      b) Al                      c) Zn                      d) None of these
242. Calamine is an ore of:  
 a) Hg                      b) Zn                      c) Cd                      d) Ca
243. The furnace which provides the highest temperature is:  
 a) Blast furnace  
 b) Reverberatory furnace  
 c) Electrical furnace  
 d) Muffle furnace
244. After partial roasting, the sulphide of copper is reduced by:  
 a) Cyanide process  
 b) Electrolysis  
 c) Reduction with carbon  
 d) Self reduction
245. Roasting is used in the extraction of:  
 a) Galena                      b) Iron pyrite                      c) Copper glance                      d) All of these
246. An ore of potassium is  
 a) Cryolite                      b) Bauxite                      c) Carnallite                      d) Dolomite
247. Metals occur in the native form because of their:  
 a) High electronegativity  
 b) High reactivity  
 c) Low reactivity  
 d) Low density

248. Purpose of smelting of an ore is  
 a) To oxidize it  
 c) To reduce it  
 b) To remove vaporisation impurities  
 d) To obtain an alloy
249. Oxidation method is used for refining of:  
 a) Pb  
 b) Cu  
 c) Hg  
 d) All of these
250. From which form of iron, other forms of iron can be produced?  
 a) Cast iron  
 b) Wrought iron  
 c) Pig iron  
 d) Steel
251. Aluminium is extracted by the electrolysis of:  
 a) Bauxite  
 b) Alumina  
 c) Molten cryolite  
 d) Alumina mixed with cryolite
252. The most abundant element in the earth crust is:  
 a) O  
 b) Si  
 c) H  
 d) C
253. Among the following statements, the incorrect one is  
 a) Calamine and siderite are carbonates  
 b) Malachite and azurite are ores of copper  
 c) Argentite and cuprite are oxides  
 d) Zinc blende and pyrites are sulphides
254. Roasting is generally carried out in case of:  
 a) Oxide ores  
 b) Sulphide ores  
 c) Silicate ores  
 d) Carbonate ores
255. Chile saltpetre is the ore of:  
 a) Mg  
 b) K  
 c) Na  
 d) Ca
256. Nickel is purified by thermal decomposition of its:  
 a) Hydride  
 b) Chloride  
 c) Azide  
 d) Carbonyl
257. Which element occurs freely in nature?  
 a) Iodine  
 b) Sulphur  
 c) Phosphorus  
 d) Magnesium
258. To dissolve argentite ore which of the following is used?  
 a)  $\text{Na}[\text{Ag}(\text{CN})_2]$   
 b)  $\text{NaCN}$   
 c)  $\text{NaCl}$   
 d)  $\text{HCl}$
259. The metal used in storage batteries is:  
 a) Cu  
 b) Sn  
 c) Pb  
 d) Ni
260. The process of Zinc -plating on iron sheet is known as  
 a) Annealing  
 b) Roasting  
 c) Galvanization  
 d) smelting
261. Bronze is a mixture of  
 a) Pb+ Sn  
 b) Cu+ Sn  
 c) Cu+ Zn  
 d) Pb+ Zn
262. Electrolytic reduction of alumina to aluminium by Hall-Heroult process is carried out  
 a) In the presence of  $\text{NaCl}$   
 b) In the presence of fluorite  
 c) In the presence of cryolite which forms a melt with lower melting point  
 d) In the presence of cryolite which forms a melt with high melting point
263. Bauxite ore is made up of  $\text{Al}_2\text{O}_3 + \text{SiO}_2 + \text{TiO}_2 + \text{Fe}_2\text{O}_3$ . This ore is treated with conc  $\text{NaOH}$  solution at 500 K and 35 bar pressure for few hours and filtered when hot. In the filtrate, the species present are  
 a)  $\text{NaAl}(\text{OH})_4$  only  
 b)  $\text{Na}_2\text{Ti}(\text{OH})_6$  only  
 c)  $\text{NaAl}(\text{OH})_4$  and  $\text{Na}_2\text{SiO}_3$  both  
 d)  $\text{Na}_2\text{SiO}_3$  only
264. In India thorium deposits are found mainly in the forms of:  
 a) Lignite  
 b) Rutile  
 c) Monazite  
 d) None
265. The luster of a metal is due to  
 a) Its high polishing  
 b) Its high density  
 c) Its chemical inertness  
 d) Presence of free electrons
266. Which is the salt of an organic acid?  
 a) Rochelle salt  
 b) Microcosmic salt  
 c) Mohr's salt  
 d) Glauber's salt
267. An element A dissolves both in acid and alkali. It is an example of:



- a) Allotropic nature of A  
 b) Dimorphic nature of A  
 c) Amorphous nature of A  
 d) Amphoteric nature of A
268. Which of the following statements about the advantages of roasting of sulphide ore before reduction is not true?  
 a)  $\Delta G_f^\circ$  of the sulphide is greater than  $\text{CS}_2$  and  $\text{H}_2\text{S}$   
 b)  $\Delta G_f^\circ$  is negative for roasting of sulphide ore to oxide  
 c) Roasting of the sulphide to oxide is thermodynamically feasible  
 d) Carbon and hydrogen are suitable reducing agents for metal sulphides
269. Which radioactive element is more abundant in India?  
 a) Thorium                      b) Uranium                      c) Radium                      d) Radon
270. Which ore contain both iron and copper?  
 a) Cuprite                      b) Chalcocite                      c) Chalcopyrite                      d) malachite
271. Galena is an ore of:  
 a) Zn                      b) Pb                      c) Sn                      d) Ca
272. The process of extraction of sodium on a commercial scale by the electrolysis of fused sodium chloride is called:  
 a) Down's process                      b) Solvay process                      c) Nelson process                      d) Castner process
273. Before introducing  $\text{FeO}$  in blast furnace, it is converted to  $\text{Fe}_2\text{O}_3$  by roasting so that:  
 a) It may not be removed as slag with silica  
 b) It may not evaporate in the furnace  
 c) Presence of it may increase the m. p. of charge  
 d) None of the above
274. Which method of purification is represented by the following equation?  

$$\text{Ti}(s) + 2\text{I}_2(g) \xrightarrow{523\text{K}} \text{TiI}_4(g) \xrightarrow{1700\text{K}} \text{Ti}(s) + 2\text{I}_2(g)$$
 a) Cupellation                      b) Poling                      c) Van Arkel                      d) Zone refining
275. Diaspore is:  
 a)  $\text{Al}_2\text{O}_3 \cdot \text{H}_2\text{O}$                       b)  $\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$                       c)  $\text{Al}_2\text{O}_3$                       d)  $\text{Al}_2\text{O}_3 \cdot 3\text{H}_2\text{O}$
276. Formula for agate is  
 a)  $\text{Na}_2\text{SiO}_3$                       b)  $\text{K}_2\text{O} \cdot \text{SiO}_2 \cdot \text{Al}_2\text{O}_3$                       c)  $\text{SiO}_2$                       d)  $\text{CaF}_2$
277. Spelter is:  
 a) Impure zinc                      b) Impure iron                      c) Pure zinc                      d) Impure Al
278. Chloride ore among the following is:  
 a) Malachite                      b) Magnesite                      c) Magnetite                      d) Rock salt
279. Magnetic separation is used for increasing concentration of the following  
 a) Calcite                      b) Horn silver                      c) Magnesite                      d) Haematite
280. Ore pitch blende is main source of  
 a) Ra                      b) Th                      c) Mg                      d) Ce
281. Which one of the following is a mineral of iron?  
 a) Pyrolusite                      b) Magnetite                      c) Malachite                      d) Cassiterite
282. Metal which can be extracted from all the three dolomite, magnesite and carnallite is  
 a) Na                      b) K                      c) Mg                      d) Ca
283. A metal which is refined by poling is  
 a) Silver                      b) Sodium                      c) Blister copper                      d) Zinc
284. The process of converting hydrated alumina into anhydrous alumina is called:  
 a) Roasting                      b) Smelting                      c) Dressing                      d) Calcination
285. Sulphide ore is:  
 a) Copper pyrites                      b) Malachite                      c) Carnallite                      d) Magnetite





286. Which metal is sometimes found in native state?  
 a) Al                                      b) Cu                                      c) Fe                                      d) Mg
287. In metallurgical process, the flux used for removing acidic impurities is:  
 a) Silica                                      b) Sodium chloride                      c) Lime stone                              d) Sodium carbonate
288. Which of the following is not an ore?  
 a) Malachite                                      b) Calamine                                      c) Satellite                                      d) Cerussite
289. Thomas slag is:  
 a) Calcium silicate  
 b) Calcium phosphate  
 c) Tricalcium phosphate and calcium silicate  
 d) Calcium ammonium phosphate
290. Leaching process is used to get:  
 a) Ag                                      b) Au                                      c) Both (a) and (b)                      d) None of these
291. The mineral of copper is:  
 a) Azurite                                      b) Malachite                                      c) Copper pyrites                      d) All of these
292. In Goldschmidt aluminothermic process, thermite mixture contains:  
 a) 3 parts  $\text{Fe}_2\text{O}_3$  and 2 parts Al  
 b) 3 parts  $\text{Al}_2\text{O}_3$  and 4 parts Al  
 c) 1 part  $\text{Fe}_2\text{O}_3$  and 3 parts Al  
 d) 3 parts  $\text{Fe}_2\text{O}_3$  and 1 part Al
293. Two compounds having the same crystal structures and analogous formulae, are called:  
 a) Isomorphous                                      b) Isotopes                                      c) Isomers                                      d) Isobars
294. When a metal is to be extracted from its ore, if the gangue associated with the ore is silica, then  
 a) A basic flux is needed                      b) An acidic flux is needed  
 c) Both basic and acidic flux are needed      d) Neither of them is needed
295. Blister copper is obtained by:  
 a) Bessemerisation                      b) Roasting                                      c) Poling                                      d) Refining
296. Which is not an ore of magnesium?  
 a) Carnallite                                      b) Dolomite                                      c) Gypsum                                      d) Magnesite
297. Which of the following metal is sometimes found native in nature?  
 a) Mg                                      b) Cu                                      c) Al                                      d) Fe
298. Match list I with List II and select the correct answer using the codes given below the lists

	List I (Types of ore)		List II (example)
1.	Oxide ore	A.	Feldspar
2.	Sulphide ore	B.	Barytes
3.	Sulphate ore	C.	Fluorspar
4.	Halide ore	D.	Galena
		E.	Corundum

- a) 1-A, 2-E, 3-B,4-C                      b) 1-B,2-D,3-C,4-A                      c) 1-B,2-D,3-E,4-A                      d) 1-E, 2-D, 3-B,4-C
299. To obtain chromium from chromic oxide ( $\text{Cr}_2\text{O}_3$ ), the method used is:  
 a) Carbon reduction  
 b) Carbon monoxide reduction  
 c) Alumino thermic process  
 d) Electrolytic reduction
300. In order to refine blister copper, it is melted in a furnace and is stirred with green logs of wood. The purpose is  
 a) To expel the dissolved gases in blister copper  
 b) To bring the impurities to surface and oxidize them  
 c) To increase the carbon content of copper  
 d) To reduce the metallic oxide impurities with hydrocarbon gases liberated from the wood



301. Hydrometallurgy is useful in the extraction of:  
 a) Sn                                      b) Al                                      c) Hg                                      d) Ag
302. Which is not an iron ore?  
 a) Haematite                              b) Limonite                              c) Cassiterite                              d) Magnetite
303. In the modern blast furnaces, the charge consists of a mixture of  
 a) Iron pyrites + bituminous coal                              b) Hydrated iron oxides + dolomite + coke  
 c) Calcined iron oxides + limestone + coke                              d) Calcined iron oxides + lime + anthracite coal
304. A substance which reacts with gangue to form fusible material is called  
 a) Flux                                      b) Slag                                      c) Catalyst                                      d) Ore
305. Which process is not used in purification of bauxite?  
 a) Hall's method                              b) Baeyer's method                              c) Serpek's method                              d) Frankland's method
306. Gallium arsenide is purified by  
 a) van-Arkel method                                      b) Zone-refining method  
 c) Electrolytic method                                      d) Liquefaction
307. Which metal is not silver white?  
 a) Ni                                      b) Cu                                      c) Na                                      d) Sn
308. In the reverberatory furnace:  
 a) The flames do not come in contact with the charge  
 b) The flames come in contact with the charge  
 c) Only hot gases come in contact with the charge  
 d) The flames are not there at all
309. Silicon is the main constituent of:  
 a) Rocks                                      b) Alloys                                      c) Animals                                      d) Plants
310. The grey cast iron contains:  
 a) Iron carbide                              b) Silicon carbide                              c) Silicon dioxide                              d) Graphite
311. Carnallite is a mineral of  
 a) Na                                      b) Zn                                      c) Cd                                      d) Mg
312. In the extraction of lead from its ore galena, an important element recovered is:  
 a) Au                                      b) Ag                                      c) Cr                                      d) C
313. Chile salt petre is an ore of  
 a) Magnesium                              b) Bromine                              c) Sodium                                      d) Iodine
314. Native silver metal forms a water soluble complex with a dilute aqueous solution of NaCN in the presence of  
 a) Nitrogen                                      b) Oxygen                                      c) Carbon dioxide                                      d) argon
315. All ores are minerals, while all minerals are not ores because  
 a) Minerals are complex compounds  
 b) The minerals are obtained from mines  
 c) The metal cannot be extracted economically from all the minerals  
 d) All of the above are correct
316. The non-fusible impurities of ores are removed by adding:  
 a) Flux                                      b) Slag                                      c) Gangue                                      d) None of these
317. Tin is extracted from tin stone by heating it in a furnace with:  
 a)  $\text{CaCO}_3$                                       b) CaO                                      c) Steam                                      d) Coal
318. Wolframite ore contains:  
 a) Zn                                      b) W                                      c) Hf                                      d) Au
319. In Serpek's process, by product obtained in the purification of bauxite is:  
 a)  $\text{Al}_2\text{O}_3$                                       b)  $\text{N}_2$                                       c)  $\text{NH}_3$                                       d) None
320. Copper can be extracted from  
 a) Dolomite                                      b) Malachite                                      c) Galena                                      d) Kupfer nickel
321. Which element is purified by Zone refining?

- a) Ge                                      b) Ge and Si                                      c) Si                                      d) None of these
322. An important characteristic property of metals is:  
 a) Their hardness  
 b) Their ability to conduct electricity  
 c) To form oxides  
 d) The stability of their compounds
323. Crystalline metal can be transformed into metallic glass by:  
 a) Alloying  
 b) Pressing into thin plates  
 c) Slow cooling of molten metal  
 d) Very rapid cooling of a spray of the molten metal
324. Metallurgy is the process of:  
 a) Concentrating the ore    b) Roasting the ore                                      c) Extracting the metal from the ore                                      d) Adding carbon to the ore in blast furnace
325. The substance not likely to contain  $\text{CaCO}_3$  is:  
 a) Sea shells                                      b) Dolomite                                      c) Marble statue                                      d) Calcined gypsum
326. In the formation of  $\text{Al}_2\text{O}_3$  large amount of heat is evolved. This property is used in:  
 a) Deoxidation                                      b) Confectionary                                      c) Indoor photography                                      d) Thermite welding
327. CO is used in the metallurgy of:  
 a) Cu                                      b) Ni                                      c) Cr                                      d) Pt
328. The electrolytic method of reduction is employed for the preparation of metals that  
 a) Are strongly electropositive                                      b) Are weakly electropositive  
 c) Are moderately electropositive                                      d) From oxides
329. Which substance can be used for drying gases?  
 a)  $\text{CaCO}_3$                                       b)  $\text{Na}_2\text{CO}_3$                                       c)  $\text{CaHCO}_3$                                       d) CaO
330. The slag obtained during the extraction of copper from copper pyrites is composed of  
 a)  $\text{Cu}_2\text{S}$                                       b)  $\text{SiO}_2$                                       c)  $\text{CuSiO}_3$                                       d)  $\text{FeSiO}_3$
331. Matte contains mainly  
 a)  $\text{Cu}_2\text{S}$  and FeS                                      b)  $\text{Cu}_2\text{S}$                                       c)  $\text{CuS}$  and  $\text{Fe}_2\text{S}_3$                                       d) Fe
332. Which statement is correct?  
 a) All minerals are ores  
 b) A mineral cannot be an ore  
 c) An ore cannot be a mineral  
 d) All ores are minerals
333. The phenomenon of removing layers of basic oxides from metals before electroplating is called:  
 a) Galvanising                                      b) Anodising                                      c) Pickling                                      d) Poling
334. Radium is obtained from:  
 a) Pitch blende                                      b) Haematite                                      c) Monazite                                      d) None of these
335. Main function of roasting is  
 a) Oxidation                                      b) Reduction  
 c) Slag formation                                      d) To remove volatile substance
336. Zinc metal is refined by:  
 a) Crystallisation                                      b) Sublimation                                      c) Heating                                      d) Distillation
337. Rutile is an ore of:  
 a) Ti                                      b) Zr                                      c) Mn                                      d) V
338. The incorrect statement is:  
 a) Calamine and siderite are carbonates  
 b) Argentite and cuprites are oxides  
 c) Zinc blende and iron pyrites are sulphides  
 d) Malachite and azurite are ores of Cu



339. Electrometallurgical process (electrolysis of fused salt) is employed to extract:  
 a) Iron                                      b) Lead                                      c) Sodium                                      d) Silver
340. Which of the following is correct?  
 a) Tin stone is magnetic in nature                                      b) Wolframite is non-magnetic in nature  
 c) Wolframeite is  $\text{FeWO}_4$                                       d) Cassiterite and rutile are sulphides ore
341. Which substance is used as basic refractory material in furnace?  
 a)  $\text{Al}_2\text{O}_3$                                       b)  $\text{SiO}_2$                                       c)  $\text{CaO}$                                       d)  $\text{Fe}_2\text{O}_3$
342. Cinnabar is:  
 a)  $\text{CuS}$                                       b)  $\text{Ag}_2\text{S}$                                       c)  $\text{ZnS}$                                       d)  $\text{HgS}$
343. Metal occur in the native form because of their  
 a) High electronegativity                                      b) High reactivity  
 c) Low reactivity                                      d) Low density
344. The method of concentrating the ore which makes use of the difference in density between ore and impurities is called  
 a) Leaching                                      b) Liquation                                      c) Levigation                                      d) Magnetic separation
345. The reaction  $2\text{ZnS} + 3\text{O}_2 \rightarrow 2\text{ZnO} + 2\text{SO}_2$  in the metallurgical process of zinc is called  
 a) Roasting                                      b) Smelting                                      c) Cupellation                                      d) Calcinations
346. In electro-refining of metal the impure metal is made the anode and a strip of pure metal, the cathode, during the electrolysis of an aqueous solution of a complex metal salt. This method cannot be used for refining of:  
 a) Silver                                      b) Copper                                      c) Aluminium                                      d) Sodium
347. Which metal is extracted by electrolytic reduction method?  
 a)  $\text{Cu}$                                       b)  $\text{Al}$                                       c)  $\text{Ag}$                                       d)  $\text{Fe}$
348. The cheap and high melting point compound used in furnace lining is:  
 a)  $\text{PbO}$                                       b)  $\text{CaO}$                                       c)  $\text{HgO}$                                       d)  $\text{ZnO}$
349. In the metallurgy of iron, when  $\text{CaCO}_3$  is added to blast furnace, calcium ion appears as  
 a) Slag                                      b) Gangue                                      c)  $\text{CaO}$                                       d) Metallic  $\text{Ca}$
350. Alloys of which metal are light and strong and are used in the manufacture of aeroplanes?  
 a)  $\text{Cr}$                                       b)  $\text{Sn}$                                       c)  $\text{Fe}$                                       d)  $\text{Mg}$
351. Which of the following processes involves the roasting process?  
 a)  $\text{ZnCO}_3 \rightarrow \text{ZnO} + \text{CO}_2$   
 b)  $\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 2\text{Fe} + 3\text{CO}$   
 c)  $2\text{PbS} + 3\text{O}_2 \rightarrow 2\text{PbO} + 2\text{SO}_2$   
 d)  $\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O} \rightarrow \text{Al}_2\text{O}_3 + 2\text{H}_2\text{O}$
352. Which of the following ore is used for industrial extraction of aluminium in India?  
 a) Corundum                                      b) Cryolite                                      c) Bauxite                                      d) Kaolin
353.  $\text{Pb}$  and  $\text{Sn}$  are extracted from their chief ore by  
 a) Electrolysis and self reduction                                      b) Self reduction and electrolysis  
 c) Carbon reduction and self reduction                                      d) Self reduction and carbon reduction
354. Heating of carbonate ores to remove carbon is called as:  
 a) Roasting                                      b) Calcination                                      c) Smelting                                      d) Fluxing
355. Coating of zinc on iron objects is commonly known as:  
 a) Electroplating                                      b) Surface coating                                      c) Galvanising                                      d) Sheardising
356. The temperature of the slag zone in the metallurgy of iron using blast furnace is  
 a)  $1200 - 1500^\circ\text{C}$                                       b)  $1500 - 1600^\circ\text{C}$                                       c)  $400 - 700^\circ\text{C}$                                       d)  $800 - 1000^\circ\text{C}$
357. Sapphire is a mineral of:  
 a)  $\text{Zn}$                                       b)  $\text{Cu}$                                       c)  $\text{Hg}$                                       d)  $\text{Al}$



## GENERAL PRINCIPLES AND PROCESSES OF ISOLATION OF ELEMENTS

### : ANSWER KEY :

1) a	2) a	3) c	4) b	153) b	154) a	155) c	156) d
5) b	6) b	7) b	8) b	157) d	158) c	159) a	160) b
9) c	10) c	11) c	12) a	161) c	162) a	163) c	164) d
13) a	14) c	15) c	16) d	165) c	166) d	167) c	168) c
17) c	18) b	19) d	20) a	169) b	170) d	171) d	172) a
21) a	22) b	23) c	24) b	173) a	174) d	175) b	176) d
25) d	26) a	27) d	28) b	177) c	178) d	179) c	180) a
29) a	30) b	31) c	32) c	181) c	182) c	183) c	184) b
33) d	34) a	35) a	36) d	185) d	186) b	187) a	188) b
37) a	38) c	39) b	40) a	189) c	190) b	191) c	192) c
41) a	42) c	43) a	44) d	193) c	194) b	195) d	196) c
45) b	46) a	47) c	48) d	197) c	198) b	199) a	200) a
49) c	50) b	51) c	52) b	201) b	202) a	203) d	204) a
53) c	54) a	55) b	56) a	205) c	206) c	207) c	208) d
57) b	58) d	59) d	60) c	209) d	210) b	211) b	212) b
61) c	62) a	63) b	64) a	213) b	214) b	215) a	216) b
65) b	66) a	67) b	68) a	217) b	218) c	219) b	220) d
69) a	70) c	71) c	72) c	221) d	222) b	223) b	224) c
73) c	74) a	75) c	76) a	225) d	226) c	227) d	228) c
77) a	78) b	79) c	80) a	229) c	230) c	231) b	232) d
81) b	82) c	83) c	84) b	233) b	234) a	235) b	236) d
85) b	86) c	87) a	88) c	237) c	238) d	239) a	240) c
89) b	90) d	91) d	92) b	241) a	242) b	243) c	244) d
93) b	94) b	95) c	96) c	245) d	246) c	247) c	248) c
97) b	98) c	99) d	100) d	249) d	250) c	251) d	252) a
101) d	102) a	103) b	104) c	253) c	254) b	255) c	256) d
105) b	106) b	107) a	108) b	257) b	258) b	259) c	260) c
109) a	110) d	111) a	112) a	261) b	262) c	263) a	264) c
113) a	114) a	115) c	116) a	265) d	266) a	267) d	268) d
117) d	118) d	119) d	120) d	269) a	270) c	271) b	272) a
121) b	122) b	123) a	124) b	273) a	274) c	275) a	276) c
125) a	126) a	127) a	128) a	277) a	278) d	279) d	280) a
129) b	130) b	131) c	132) d	281) b	282) c	283) c	284) d
133) a	134) d	135) d	136) d	285) a	286) c	287) c	288) c
137) c	138) d	139) b	140) a	289) c	290) c	291) d	292) d
141) d	142) a	143) c	144) d	293) a	294) a	295) a	296) c
145) b	146) b	147) b	148) a	297) b	298) d	299) c	300) d
149) c	150) d	151) b	152) a	301) d	302) c	303) c	304) a



305) d	306) b	307) b	308) c	337) a	338) b	339) c	340) c
309) a	310) d	311) d	312) b	341) c	342) d	343) c	344) c
313) c	314) b	315) c	316) a	345) a	346) d	347) b	348) b
317) d	318) b	319) c	320) b	349) a	350) d	351) c	352) c
321) b	322) b	323) d	324) c	353) d	354) b	355) c	356) d
325) d	326) d	327) b	328) a	357) d			
329) d	330) d	331) a	332) d				
333) c	334) a	335) d	336) d				



# GENERAL PRINCIPLES AND PROCESSES OF ISOLATION OF ELEMENTS

## : HINTS AND SOLUTIONS :

- |   |  |
|---|--|
| <p>1 <b>(a)</b><br/> <math display="block">\text{Ti} + 2\text{I}_2 \xrightarrow{500\text{ K}} \underset{\substack{\text{volatile} \\ \text{stable compound}}}{\text{TiI}_4} + 2\text{I}_2 \xrightarrow{1700\text{ K}} \text{Ti} \text{ pure metal}</math></p> <p>2 <b>(a)</b><br/>                 Cryolite has these two functions during electrolysis of alumina.</p> <p>3 <b>(c)</b><br/>                 Pyrolusite is an ore of Mn containing <math>\text{MnO}_2</math>.</p> <p>4 <b>(b)</b><br/>                 Wrought or malleable iron is the purest form of iron</p> <p>5 <b>(b)</b><br/>                 ___do___</p> <p>6 <b>(b)</b><br/>                 During electrolysis, noble metals (inert metals) like Ag, Au and Pt are not affected and separate as anode mud from the impure anode</p> <p>7 <b>(b)</b><br/> <math>\text{PbS} + 2\text{PbO} \rightarrow 3\text{Pb} + \text{SO}_2</math> (Self reduction)<br/> <math>\text{SnO} + \text{C} \rightarrow \text{Sn} + \text{CO}</math> (Carbon reduction)</p> <p>8 <b>(b)</b><br/>                 At about 330 K nickel is attacked by carbon monoxide with the formation of a volatile nickel carbonyl <math>\text{Ni}(\text{CO})_4</math>.</p> <p>10 <b>(c)</b><br/>                 It involves auto-reduction.<br/> <math display="block">2\text{Cu}_2\text{O} + \text{Cu}_2\text{S} \rightarrow 6\text{Cu} + \text{SO}_2</math></p> <p>11 <b>(c)</b><br/>                 In electrolytic refining of Cu, impurities of Fe, Ni, and Zn pass into solution and others like Au and Ag fall down, as anode mud.</p> <p>12 <b>(a)</b><br/>                 Mercury is the only metal which is liquid at room temperature.</p> <p>13 <b>(a)</b><br/>                 Lapis lazuli is the sodium alumino silicate present in earth rocks as blue stone</p> <p>14 <b>(c)</b><br/> <math>2\text{MS} + \text{C} \rightarrow 2\text{M} + \text{CS}_2 \quad \Delta G_1 = \text{positive}</math><br/> <math>2\text{MO} + \text{C} \rightarrow 2\text{M} + \text{CO}_2 \quad \Delta G_2 = \text{negative}</math><br/>                 The value of <math>\Delta G</math> for the formation of <math>\text{CO}_2</math> is negative, <i>ie</i>, it is thermodynamically more than</p> | <p><math>\text{CS}_2</math>. Also metal sulphides are thermodynamically more stable than <math>\text{CS}_2</math>. Metal sulphides are more stable than the corresponding oxides, so they are roasted to convert into less stable oxides</p> <p>15 <b>(c)</b><br/>                 It is a fact.</p> <p>16 <b>(d)</b><br/>                 Soda ash (<math>\text{Na}_2\text{CO}_3</math>) is an ore of sodium</p> <p>17 <b>(c)</b><br/>                 Titanium is quite abundant in nature and mainly occurs as ilmenite, <math>\text{FeO} \cdot \text{TiO}_2</math></p> <p>20 <b>(a)</b><br/>                 It is a fact.</p> <p>21 <b>(a)</b><br/>                 It is a fact.</p> <p>22 <b>(b)</b><br/>                 Al is highly electropositive. It can be obtained only by electrolytic reduction</p> <p>23 <b>(c)</b><br/>                 Malleable nature (<i>ie.</i>, can be pressed out into sheets) is maximum in gold.</p> <p>24 <b>(b)</b><br/>                 It is a fact.</p> <p>25 <b>(d)</b><br/>                 It is a fact.</p> <p>26 <b>(a)</b><br/>                 Cassiterite is an ore of tin</p> <p>27 <b>(d)</b><br/> <math>2\text{PbS} + 3\text{O}_2 \rightarrow 2\text{PbO} + 2\text{SO}_2</math><br/> <math>\text{PbS} + 2\text{PbO} \rightarrow 3\text{Pb} + \text{SO}_2</math></p> <p>28 <b>(b)</b><br/>                 Cryolite is an ore of Al containing <math>\text{Na}_3\text{AlF}_6</math>.</p> <p>29 <b>(a)</b><br/>                 Blanc fixe is <math>\text{BaSO}_4</math>.</p> <p>30 <b>(b)</b><br/>                 Froth floatation method is based on the fact that the surface of sulphide ores is preferentially wetted by oil while that of gangue is wetted by water</p> <p>31 <b>(c)</b><br/>                 Fe-C form alloy.</p> <p>32 <b>(c)</b></p> |
|---|--|

- Galena is PbS; Sulphide ores are concentrated by froth floatation process.
- 33 (d)  
In Hall and Heroult process,  

$$2\text{Al}_2\text{O}_3 \rightarrow 4\text{Al} + 3\text{O}_2$$

$$4\text{C} + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{CO} \uparrow$$

$$2\text{Al}_2\text{O}_3 + 4\text{C} \rightarrow 4\text{Al} + 2\text{CO}_2 + 2\text{CO}$$
Only for removal of  $\text{CO}_2$ , following equation is possible  

$$2\text{Al}_2\text{O}_3 + 3\text{C} \rightarrow 4\text{Al} + 3\text{CO}_2$$

$$3 \times 12 = 36 \quad 4 \times 27 = 108$$
 $\therefore$  For 108 g of Al, required amount of C = 36g  
 $\therefore$  For 270 g of required amount of C =  $\frac{36}{108} \times 270 = 90\text{g}$
- 34 (a)  

$$\text{CuO} + \text{CO} \xrightarrow{\Delta} \text{Cu} + \text{CO}_2 \uparrow$$
- 35 (a)  
Load stone (magnetite,  $\text{Fe}_3\text{O}_4$ ) is an ore of iron
- 36 (d)  
Mond's process for refining of Ni is an example of vapour phase refining
- 37 (a)  
Carbon reduction process is used for extraction of less electropositive metals like Pb, Fe, Zn, Sb, Cu, etc., from their ores.
- 38 (c)  
The phenomenon of efflorescence involves spontaneous loss of water molecules from a crystal.
- 39 (b)  
Cd is found as traces in most Zn ores, and is extracted from these.  

$$\text{Zn}_{(\text{solid})} + \text{Cd}_{(\text{solution})}^{2+} \rightarrow \text{Zn}_{(\text{solution})}^{2+} + \text{Cd}_{(\text{solid})}; E^\ominus = 0.36\text{ V}$$
- 41 (a)  

$$2\text{MnO}_2 + 4\text{KOH} + \text{O}_2 \rightarrow 2\text{K}_2\text{MnO}_4 + 2\text{H}_2\text{O}$$
Purple green
- 42 (c)  
 $\text{SiO}_2$  is an acidic flux.
- 43 (a)  
 $\text{FeCrO}_4$  is magnetic impurity.
- 44 (d)  
Extraction of Ni involves Electrolytic Process, Oxford Process, Mond's Process and German Process.4
- 45 (b)  
Acetate of all metals are soluble in water.
- 46 (a)  
\_\_\_\_do\_\_\_\_
- 47 (c)  
Electromagnetic separation is used when either the ore or the impurities associated with it, are magnetic in nature
- 48 (d)  
Alkaline earth metals are very reactive and are found in combined state only in nature.
- 49 (c)  
Dispersion of solid in solid is called solid sol.
- 51 (c)  
Pt is noble metal, other noble metals are Au, Ag.
- 52 (b)  
Alumino-thermic process is commonly used for those metals which have very high m.pt. and are to be extracted from their oxides and their reduction with carbon is not satisfactory.
- 53 (c)  
Bauxite ore is concentrated by chemical separation or leaching. In this, powdered ore is treated with a suitable reagent which can dissolve the ore but not the impurities
- 54 (a)  
Dressing or beneficiation of ore involves removal of impurities from ore.
- 55 (b)  
Zinc blende is roasted and then treated with coke for the reduction  

$$3\text{ZnS} + 3\text{O}_2 \xrightarrow{\Delta} 2\text{ZnO} + 2\text{SO}_2 \uparrow$$

$$\text{ZnO} + \text{C} \xrightarrow{\Delta} \text{Zn} + \text{CO} \uparrow$$
- 56 (a)  

$$2\text{HgS} + 3\text{O}_2 \rightarrow 2\text{HgO} + 3\text{SO}_2$$

$$2\text{HgO} \xrightarrow{\Delta} 2\text{Hg} + \text{O}_2$$

$$2\text{Cu}_2\text{S} + 3\text{O}_2 \rightarrow 2\text{Cu}_2\text{O} + 2\text{SO}_2$$

$$2\text{Cu}_2\text{O} + \text{Cu}_2\text{S} \rightarrow 6\text{Cu} + \text{SO}_2$$
- 57 (b)  
Mixture of calcium phosphate and calcium silicate is known as Thomas slag
- 58 (d)  
Metals are good conductor of electricity because they contain free electrons
- 59 (d)  
Purification of Hg, Sn and Bi involves liquation.
- 60 (c)  
The abundance of elements in earth crust follow the order  $\text{O} > \text{Si} > \text{Al} > \text{Fe}$ .
- 61 (c)  
Sulphur exists in various allotropic forms such as rhombic, monoclinic, plastic forms.
- 62 (a)



- Van-Arkel method is not used for extraction of Al. it is used in the purification of Ti
- 63 **(b)**  
Indian saltpetre is a nitrate ore of K containing  $\text{KNO}_3$ .
- 64 **(a)**  
It is a fact.
- 65 **(b)**  
It is an ore of Pt.
- 66 **(a)**  
Pine oil is foaming agent. An another substance called collector such as potassium ethyl xanthate or amyl xanthate is also added
- 67 **(b)**  
It is a fact.
- 68 **(a)**  
Lime stone is used as basic flux to fuse acidic impurities of Silica.  

$$\text{CaCO}_3 + \text{SiO}_2 \rightarrow \text{CaSiO}_3 + \text{CO}_2 \uparrow$$
Flux      Gangue      Slag
- 69 **(a)**  
Auto reduction occurs  

$$\text{Cu}_2\text{S} + 2\text{Cu}_2\text{O} \rightarrow 6\text{Cu} + \text{SO}_2$$
- 70 **(c)**  
Cassiterite is an ore of tin
- 71 **(c)**  

$$\text{ZnO} + \text{CO} \rightarrow \text{CO}_2 + \text{Zn}$$
- 72 **(c)**  
Zinc blende is an ore of Zn containing ZnS.
- 73 **(c)**  
Hg having low b.pt. is easily distilled off.
- 74 **(a)**  
Density increases with increasing atomic number.
- 75 **(c)**  
Calamine is the carbonate ore zinc ( $\text{ZnCO}_3$ )
- 76 **(a)**  

$$\text{CaCO}_3 + \text{SiO}_2 \rightarrow \text{CaSiO}_3$$
Slag
- 77 **(a)**  
Mond's process is used for the purification of Ni  

$$\text{Ni} + 4\text{CO} \xrightarrow{\text{Heat}} [\text{Ni}(\text{CO})_4] \xrightarrow{\text{Decompose}} \text{Ni} + 4\text{CO}$$
- 78 **(b)**  
Dolomite is an ore containing,  $\text{CaCO}_3 \cdot \text{MgCO}_3$ .
- 79 **(c)**  
\_\_do\_\_
- 80 **(a)**  
It is a fact.
- 81 **(b)**  
Zinc blende is ZnS not  $\text{ZnCl}_2$
- 82 **(c)**
- Mg and Al cannot be obtained by the electrolysis of aqueous solution of their salts because instead of metal,  $\text{H}_2$  gas is liberated at cathode
- 83 **(c)**  
Lime stone acts as basic flux for sandstone ( $\text{SiO}_2$ ).
- 85 **(b)**  
The main impurity in red bauxite is ferrite ( $\text{Fe}_2\text{O}_3$ ) and the main impurity in white bauxite is silica ( $\text{SiO}_2$ )
- 86 **(c)**  
Rutile is  $\text{TiO}_2$ .
- 88 **(c)**  

$$\text{FeO} + \text{CO} \xrightarrow{1000^\circ\text{C}} \text{Fe} + \text{CO}_2$$
- 89 **(b)**  
Silver is recovered from the alloy (lead-silver alloy) by cupellation.
- 90 **(d)**  
Hg has low b. pt. and is purified by distillation.
- 91 **(d)**  
Lepidolite is  $(\text{Li, Na, K})_2; \text{Al}_2(\text{SiO}_3)_3, (\text{F, OH})_2$ .
- 92 **(b)**  
VIA group member or oxygen family is known as chalcogens.
- 93 **(b)**  

$$\text{Al}_2\text{O}_3 + \text{Na}_2\text{CO}_3 \rightarrow 2\text{NaAlO}_2 + \text{CO}_2$$
- 94 **(b)**  
Stibnite is an ore of Sb containing  $\text{Sb}_2\text{S}_3$ .
- 96 **(c)**  
Baryte is an ore of Barium having formula  $\text{BaSO}_4$ .
- 97 **(b)**  
Thermite is a mixture of Al and  $\text{F}_2\text{O}_3$  in 1:3 ratio
- 98 **(c)**  
It is definiton of calcination.
- 99 **(d)**  
Following reaction takes place during bessemerisation  

$$2\text{Cu}_2\text{O} + \text{Cu}_2\text{S} \rightarrow 6\text{Cu} + \text{SO}_2$$
- 101 **(d)**  
Calcination involves decomposition of ore to remove volatile impurities.
- 102 **(a)**  

$$\text{FeO} + \text{SiO}_2 \rightarrow \text{FeSiO}_3 \text{ (Fusible slag)}$$

$$\text{Cu}_2\text{O} + \text{FeS} \rightarrow \text{Cu}_2\text{S} + \text{FeO}$$
Slag is removed from the slag hole while a molten mass containing mostly cuprous sulphide with a little ferrous sulphide called matte.
- 103 **(b)**  
Mond's process is used for the purification of Ni.
- 104 **(c)**  
It is a fact.

- 105 **(b)**  
Bauxite ( $\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$ )  
Corundum ( $\text{Al}_2\text{O}_3$ )  
Diaspore ( $\text{Al}_2\text{O}_3 \cdot \text{H}_2\text{O}$ )
- 106 **(b)**  
Extraction of less electropositive metals say Cr, Mn, Cu, Ca, Ni, etc., can be done by heating their oxides with strong reducing agents, e.g.,  $\text{CO}$ ,  $\text{CO} + \text{H} + \text{Na}$ , Al, Mg, etc.
- 107 **(a)**  
Lead extracted from argentiferrous galena contains small quantities of silver. Recovery of silver from argentiferrous lead is an economical proposition and is carried out by Parke's process.
- 108 **(b)**  
Oxygen family is known as chalcogens.
- 109 **(a)**  
Apatite is  $\text{CaF}_2 \cdot 3\text{Ca}_3(\text{PO}_4)_2$   
 $\therefore$  It is ore of fluorine with calcium
- 110 **(d)**  
A bronze sulphide mineral  $(\text{Fe}, \text{Ni})_9\text{S}_8$ , a chief ore c
- 111 **(a)**  
Pitch blende is an ore of uranium containing  $\text{U}_3\text{O}_8$ .
- 112 **(a)**  
In alumino-thermic process, aluminium is used as reducing agent
- 113 **(a)**  
\_\_do\_\_
- 114 **(a)**  
The adsorption phenomenon is involved in the forth floatation process
- 115 **(c)**  
It is a fact.
- 116 **(a)**  
Stainless steel is an alloy of iron with chromium and niel Its composition is 82% Fe and 18% (Cr+Ni). It res corrosion and used for making automobile parts and utensils
- 117 **(d)**  
It is a fact.
- 118 **(d)**  
Cyanide process is used in the extraction of both silver and gold because these form complex salts with  $\text{CN}^-$  ion due to presence of lone pair of electron on nitrogen atom
- 119 **(d)**  
All are characteristic features of alloy.
- 120 **(d)**  
All are noble metals.
- 121 **(b)**  
Metals which are inert towards many common reagents are called noble metals.
- 122 **(b)**  
Follow text.
- 123 **(a)**  
Fe does not form amalgam with Hg.
- 124 **(b)**  
Peat is an early stage in the formation of coal from vegetable matter.
- 125 **(a)**  
It is a fact.
- 126 **(a)**  
It is definition of roasting.
- 127 **(a)**  
Due to the formation of  $\text{K}_2\text{MnO}_4$ .
- 128 **(a)**  
Other methods are used for extraction of Al from its ores.
- 129 **(b)**  
It is a fact.
- 130 **(b)**  
The method of zone refining of metals is based on the principle of greater solubility of the impurity in the molten state than in the solid. Elements which are used as semiconductors like Si, Ge, Ga etc are refined by this method
- 131 **(c)**  
Bauxite is  $\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$ .
- 132 **(d)**  
Pyrolusite –  $\text{MnO}_2$   
Malachite –  $\text{CuCO}_3 \cdot \text{Cu}(\text{OH})_2$   
Diaspore –  $\text{Al}_2\text{O}_3 \cdot \text{H}_2\text{O}$   
Cassiterite –  $\text{SnO}_2$
- 133 **(a)**  
Fluorspar ( $\text{CaF}_2$ ),  
Cryolite ( $\text{Na}_3\text{AlF}_6$ )  
Feldspar ( $\text{KAlSi}_3\text{O}_8$ ),  
Mica ( $\text{K}_2\text{O} \cdot 3\text{Al}_2\text{O}_3 \cdot 6\text{SiO}_2 \cdot 2\text{H}_2\text{O}$ )
- 134 **(d)**  
Alloys of metals with Hg are called amalgams.
- 135 **(d)**  
A natural sulphide of iron and arsenic,  $\text{FeAsS}$ .
- 136 **(d)**  
In froth floatation method, the pure ore is not easily wetted by water but wetted by pine oil, so it is successfully separated from impurities
- 137 **(c)**  
Metallic character increases down the group.
- 138 **(d)**  
Molten magnesium is lighter than ore.

- 140 (a)  
This is auto reduction of copper sulphide.
- 141 (d)  
Sulphide ores on roasting forms oxide and give  $\text{SO}_2$ .
- 142 (a)  
The compounds which combine with impurities preseore (at high temperature) and remove them as a fls substance (slag) are known as flux. When basic impurities are present, an acidic flux is used and *vice – versa*  

$$\text{FeO} + \text{SiO}_2 \rightarrow \text{FeSiO}_3$$
Basic impurity    acidic flux    slag
- 144 (d)  
It is a fact.
- 145 (b)  
Wolframite ore [ $\text{FeWO}_4$ ] is present in tin stone as impurities and it has same mass per unit volume as that of tin stone. So, it is separated by electromagnetic separator because wolframite is magnetic in nature, hence it gets attached by magnet while tin stone does not
- 146 (b)  
Fe ores are magnetic in nature.
- 147 (b)  
Because reduction of highly electropositive elements, (e.g., alkali metals, alkaline earth metals and Al) cannot be made by other metals.
- 148 (a)  
In blast furnace, at the top is the zone of reduction. Here  $\text{Fe}_2\text{O}_3$  is reduced to spongy iron by CO rising up.  

$$\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$$
- 149 (c)  
It is a fact.
- 150 (d)  
Cyanide process is used for obtaining silver. This process is also called as Mac Arthur and Forest process
- 151 (b)  
It is a fact.
- 152 (a)  
In Bessemer converter, copper sulphide is partially oxidised to cuprous oxide which further reacts with remaining copper sulphide to form copper and sulphur dioxide.  

$$\text{Cu}_2\text{S} + 2\text{Cu}_2\text{O} \rightarrow 6\text{Cu} + \text{SO}_2$$
- 153 (b)  
Beryl is  $3\text{BeO} \cdot \text{Al}_2\text{O}_3 \cdot 6\text{SiO}_2$ .
- 154 (a)  
It is a fact.
- 155 (c)  
Lead present as impurity in the silver obtained by argentiferous lead is purified by cupellation.
- 156 (d)  
 $\text{CaO}$ ,  $\text{K}_2\text{O}$  cannot reduced by carbon reduction method
- 158 (c)  
The method is used for purification of Zr and Ti in which these metals on heating with  $\text{I}_2$  forms vapours of metal iodide which on decomposition gives pure metals.
- 159 (a)  
Mass number of uranium is highest, i.e.,  $\text{U}^{238}$ .
- 160 (b)  
Iron is made inactive or passive by oxidizing agents like conc. Nitric acid, chromic acid, acidified  $\text{KMnO}_4$ , etc., the cause of this is the formation of a thin film of oxide on the surface of the metal.
- 161 (c)  
Kiesserite is an ore of Mg containing,  $\text{MgSO}_4 \cdot \text{H}_2\text{O}$ .
- 162 (a)  
In smelting carbon is used for the reduction of oxide to metal.
- 163 (c)  
Metallic character increases down the gp.
- 164 (d)  
Ag, Au are obtained by complex formation.
- 165 (c)  
Diamond consists of carbon atoms only.
- 166 (d)  
Impurities are known as matrix or gangue.
- 167 (c)  
 $\text{CaCN}_2$  is used as a fertilizer.
- 168 (c)  
It is a fact.
- 169 (b)  
It is found in human body as haemoglobin.
- 170 (d)  
Flux is used to fuse non-fusible impurities (both acidic and basic) present in the ore
- 171 (d)  
Flux is a substance with combine with gangue that present in the roasted or calcined ore to form fusible product, called slag
- 172 (a)  
Oxides of less electropositive metals such as  $\text{Cr}_2\text{O}_3$ , Mn etc are reduced by using Al. This process is called them process

- $\text{Cr}_2\text{O}_3 + 2\text{Al} \rightarrow \text{Al}_2\text{O}_3 + 2\text{Cr} + \text{Heat}$
- 173 (a) Gangue particles are wetted up by water and adsorbed.
- 174 (d) Composition of various alloy of steels are as-Nickel steel -3.5% Ni, Chrome steel-1.5-2% Cr, Chrome-vanadium -0.15% V, 1 % Cr, Manganese steel-1.2-15% Mn, Tungsten steel 14-20% W, 3-8% Cr, Invar 36% Ni, Stainless steel 11.5% Cr.
- 175 (b) Leaching is used to make insoluble ore in soluble form.
- 176 (d) In electrorefining of copper, some gold is deposited as anode mud
- 177 (c) Thus, furnace material can withstand high temperature.
- 178 (d) Anode :  $\text{Cu} \rightarrow \text{Cu}^{2+} + 2e$   
(Impure sample)  
Cathode :  $\text{Cu}^{2+} + 2e \rightarrow \text{Cu}$   
(Pure Cu)
- 179 (c)  $\text{Ag}_2\text{S}$  forms soluble complex with KCN.
- 180 (a) Van-Arkel method is used to purify metals such as Zr, Ti, V, Th, etc, limestone is basic flux. Dolomite ( $\text{CaCO}_3$ ) is an ore of Ca. Willemite ( $\text{Zn}_2\text{SiO}_4$ ) is a silicate ore
- 181 (c) Froth-floatation is used to concentrated sulphide ores [Galena  $\text{PbS}$ ]
- 182 (c) Borax and Colemanite both are the ores of Boron containing  $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10\text{H}_2\text{O}$  and  $\text{Ca}_2\text{B}_6\text{O}_{11} \cdot 5\text{H}_2\text{O}$  respectively.
- 183 (c) Follow Mc Arthur-Forest process for Ag .
- 184 (b) S is oxidised to  $\text{SO}_2$  (g).
- 185 (d) All are magnetic ores.
- 186 (b) A method for purification of titanium metal.
- 187 (a) It is fact.
- 188 (b) Follow extraction of Zn.
- 189 (c) It is a fact.
- 190 (b) Flux is mixed with concentrated ore which is not soluble in molten metal
- 191 (c) It is a fact.
- 192 (c) It is a fact.
- 193 (c) Magnesium chloride is present in sea water.
- 194 (b) Wulfenite is a molybdate containing Pb, Mo,  $\text{O}_4$  .
- 195 (d) The fourth -floatation process is based upon the preferential wetting of ore particle by oil
- 196 (c) It is a fact.
- 197 (c) Lead dissolves in water containing dissolved air, due to the formation of lead hydroxide. This solvent action of water on lead is called plumbo solvency.
- 198 (b)  $\text{ZnO} \xrightarrow{\text{Reduction}} \text{Zn}$
- 200 (a) Refractory materials are the substances which can withstand very high temperature without melting or becoming salt
- 201 (b) It is a fact.
- 202 (a) Willemite, a rare zinc silicate mineral, is  $\text{Zn}_2\text{SiO}_4$ . It has trigonal symmetry and is strongly fluorescent green
- 203 (d)  $\text{Ag}_2\text{O}$  is decomposed on simple heating.
- 204 (a) It is a fact.
- 205 (c) \_\_\_do\_\_\_
- 206 (c) Although presence of  $\text{CO}_2$  enhances rusting due to formation of more  $\text{H}_3\text{O}^+$  ions.
- 207 (c) Ruby is mineral of aluminium, *ie*,  $\text{Al}_2\text{O}_3$ . It does not contain silicon
- 208 (d) Because all nitrates are water soluble.
- 209 (d)



- Roasting is a process in which ore is heated in air to remove Sulphur impurities.
- 210 (b) Leaching is a chemical method for the concentration of an ore.
- 212 (b) Cinnabar (HgS) is a sulphide ore, hence it is concentrated by froth floatation process
- 214 (b) Bauxite ( $\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$ ) is an oxide ore of aluminium
- 215 (a) Carnallite is an ore of magnesium containing  $\text{KCl} \cdot \text{MgCl}_2 \cdot 6\text{H}_2\text{O}$ .
- 216 (b) Siderite is  $\text{FeCO}_3$ .
- 217 (b) It is a fact.
- 218 (c) Corundum ( $\text{Al}_2\text{O}_3$ ) is the combined state of aluminium
- 219 (b) \_\_\_do\_\_\_
- 220 (d) Pig iron is the most impure form of iron and contains highest proportion of carbon (2.5-4%)  
 Malachite  $\rightarrow \text{Cu}(\text{OH})_2 \cdot \text{CuCO}_3$  (Cu ore)  
 Zinc blende  $\rightarrow \text{ZnS}$  (Zn ore)  
 Bauxite  $\rightarrow \text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$  (Al ore)
- 222 (b) Al acts as strong reducing agent and converts many metal oxides (excepts I and II gp) to metals.
- 223 (b) Cassiterite is an ore of tin.
- 224 (c) Haematite is an ore of Fe.
- 225 (d) It is an ore of Mg containing  $\text{MgCO}_3$ .
- 226 (c)
- |                 |        |
|-----------------|--------|
| Combustion zone | 1800 K |
| Fusion zone     | 1600 K |
| Slag zone       | 1300 K |
| Reduction zone  | 800 K  |
- 227 (d) All minerals are not suitable for the extraction for the extraction of metals commercially. Thus, all ores are minerals, but all minerals are not ores
- 228 (c) It is a fact.
- 229 (c) Lepidolite is  $(\text{Li, K, Na})_2 \text{Al}_2(\text{SiO}_3)_3 \cdot (\text{F} \cdot \text{OH})_2$ .
- 230 (c)  $\text{Au} + 3\text{HNO}_3 + 4\text{HCl} \rightarrow \text{HAuCl}_4 + 3\text{NO}_2 + 3\text{H}_2\text{O}$
- 231 (b) The slag float over molten mass.
- 232 (d) Cassiterite is a principal ore of tin containing  $\text{SnO}_2$ .
- 233 (b) It is a fact.
- 234 (a) Pyrolusite ( $\text{MnO}_2$ ) is not a sulphide ore, so it is not concentrated by froth floatation process
- 235 (b) Follow text.
- 236 (d) Dollucite is caesium aluminium silicate containing about 30% of caesium.
- 237 (c) Roasting is the process in which the ore is heated strongly below its melting point in presence of air
- 238 (d) Anglesite is  $\text{PbSO}_4$ .
- 239 (a) It is a fact.
- 240 (c) Pine oil reduces the surface tension of water and the solution forms froths.
- 241 (a) In smelting, powerful reducing agents like C,  $\text{H}_2$ , CO etc are used
- 242 (b) Calamine is an ore of Zn containing  $\text{ZnCO}_3$ .
- 243 (c) On striking the electric arc between the electrodes, high temperature is produced due to which the charge melts.
- 244 (d)  $2\text{CuO} + \text{CuS} \rightarrow 3\text{Cu} + \text{SO}_2 \uparrow$
- 245 (d) Roasting is mainly used in the extraction of sulphide ores. Galena— $\text{PbS}$ , Iron pyrite— $\text{FeS}$ , Copper glance— $\text{Cu}_2\text{S}$ .
- 247 (c) It is a fact.
- 248 (c) Smelting is a process of reducing metal oxide to metal by means of coke or CO  
 $\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 2\text{Fe} + 3\text{CO}$   
 $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
- 249 (d)



- Pb, Cu, and Hg all are refined by Oxidation method.
- 250 (c) Cast iron, wrought iron and steel may be produced from pig iron.
- 251 (d) Follow text.
- 252 (a) The abundance ratio :  $O > Si > Al > Fe$ .
- 253 (c) Cuprite ( $Cu_2O$ ) is oxide but argentite ( $Ag_2S$ ) is not oxide
- 254 (b) To convert ores into oxides and remove Sulphur as volatile  $SO_2$ .
- 255 (c) Chile saltpetre is  $NaNO_3$ .
- 256 (d) 
$$Ni(CO)_4 \xrightarrow{\Delta} Ni + 4CO$$
- 257 (b) Sulphur occurs in native state while iodine, phosphorus and magnesium are found in combined state.
- 258 (b) Argentite is an ore of Ag having composition  $Ag_2S$ . It dissolves in NaCN due to formation of soluble complex  

$$Ag_2S + 4NaCN \rightarrow 2Na[Ag(CN)_2] + NaCl$$
 $\therefore NaCN$  is used to dissolve argentite
- 259 (c) It is one use of lead.
- 260 (c) The process of zinc -plating on iron-sheet is known as galvanisation
- 261 (b) Bronze is mixture of Cu and Sn
- 262 (c) In the extraction of Al,  $Al_2O_3$  is melted with cryolite [ $Na_3(AlF_6)$ ]. Cryolite improves the electrical conductivity of the alumina and lowers the m.p. of the mixture to about  $950^\circ C$
- 263 (a) In bauxite ore, only  $Al_2O_3$  reacts with conc NaOH and forms sodium meta aluminate. This, further dissolves in water  

$$Al_2O_3 + 2H_2O + 2NaOH \xrightarrow[35 \text{ bar}]{500 \text{ K}} 2NaAlO_2 + 3H_2O$$

$$NaAlO_2 + 2H_2O \rightarrow 2NaAl(OH)_4$$
- 264 (c) A mineral containing phosphates oscerium, thorium and other rare earths, with some occluded helium.
- 265 (d) Luster of metals is due to the presence of mobile electrons
- 266 (a) Rochelle salt is potassium, sodium tartarate.
- 267 (d) Amphoteric compounds are soluble in both alkali and acid.
- 268 (d)  $\Delta G_f^\circ$  for sulphides  $> \Delta G_f^\circ$  of  $CS_2$  and  $H_2S$  and thus, C and  $H_2$  cannot reduce metal sulphide.
- 269 (a) It is a fact.
- 270 (c) Among cuprite [ $Cu_2O$ ], chalcocite [ $Cu_2S$ ], chalcopyrite [ $CuFeS_2$ ] and malachite [ $Cu(OH)_2 \cdot CuCO_3$ ]; only chalcopyrite is an ore which contains both Fe and Cu
- 271 (b) Galena is an ore of Pb containing  $SnO_2$ .
- 272 (a) It is a fact.
- 273 (a) FeO can form slag with  $SiO_2$ ,  

$$SiO_2 + FeO \rightarrow FeSiO_3$$
- 274 (c) This is Van-Arkel method for purification of Ti.
- 275 (a) Diaspore is an ore of aluminium containing  $Al_2O_3 \cdot H_2O$ .
- 276 (c) Quartz is found in many varieties which have different colour due to impurities, eg, amethyst (purple), opal (white) carnelian and agate  
 $\therefore$  Agate is  $SiO_2$
- 277 (a) In the metallurgy of zinc, reduction of roasted ore ( $ZnO$ ) gives impure zinc (in fire-clay retort) called spelter.
- 278 (d) Rock salt is NaCl.
- 280 (a) Ore pitch blende is main source of radium
- 281 (b) Magnetite is  $Fe_3O_4$ .
- 282 (c) Dolomite  $MgCO_3 \cdot CaCO_3$



Magnesite  $MgCO_3$   
 Carnallite  $KCl \cdot MgCl_2 \cdot 6H_2O$

283 (c) Poling is used for purification of metal which contain their own oxide as impurity, *eg*,  $Cu_2O$  in Cu;  $SnO_2$  in Sn

284 (d)  $Al_2O_3 \cdot 2H_2O \xrightarrow{-2H_2O} Al_2O_3$

The process is known as calcination, *i.e.*, to heat a mineral below its m.pt. in absence of air in order to remove moisture, organic impurities and volatile impurities.

285 (a) Copper pyrite is  $CuFeS_2$ .

286 (c) Small quantity of iron occur in native state while Al, Cu and Mg are found in combined state.

287 (c)  $SiO_2 + CaCO_3 \rightarrow CaSiO_3 + CO_2$   
 Impurity (acidic) Flux (basic) Slag

289 (c) Thomas slag is tricalcium phosphate and calcium silicate.

290 (c) Leaching process involves the treatment of the ore with a suitable reagent so as to make it soluble while impurities remain insoluble. It is used to get Ag and Au both.

291 (d) All are the mineral of copper.  
 Azurite- $Cu(OH)_2 \cdot 2CuCO_3$ , Malachite- $Cu(OH)_2 \cdot CuCO_3$ ,  
 Copper pyrites- $CuFeS_2$ .

292 (d) It is a fact.

293 (a) \_\_\_do\_\_\_

294 (a)  $SiO_2 + CaO \rightarrow CaSiO_3$   
 acidic impurity basic flux slag

295 (a) In Bessemerisation, the molten mass is run into sand moulds and allowed to solidify, when it gives out dissolved  $SO_2$  leaving blister type appearance on copper which is popularly known as blister copper.

296 (c) Gypsum is  $CaSO_4 \cdot 2H_2O$ .

297 (b)

Copper is found in native as well as in combined state

298 (d)

List I (Types of ore)	List II (Example)
Oxide ore	Corundum ( $Al_2O_3$ )
Sulphide ore	Galena (pbS)
Sulphate ore	Barytes ( $BaSO_4$ )
Halide ore	Fluorsper ( $CaF_2$ )
	Feldspar

299 (c)  $Cr_2O_3 + 2Al \rightarrow Al_2O_3 + 2Cr$ ;  $\Delta H = -ve$ .

300 (d)  $2Cu_2S + 3O_2 \rightarrow 2Cu_2O + 2SO_2$   
 $3Cu_2O + CH_4 \rightarrow 6Cu + 2H_2O + CO$   
 (from green logs of wood)

301 (d) Extraction of silver and gold is done by hydrometallurgical process or complex salt formation method.

302 (c) Cassiterite is  $SnO_2$ .

306 (b) Gallium arsenide is purified by zone refining method

307 (b) Copper metal is reddish brown in colour.

308 (c) \_\_\_do\_\_\_

309 (a) All the rocks contains silicates.

310 (d) It is a fact.

312 (b) Lead extracted from galena contains little Ag. Recovery of Ag from argentiferous lead is made by Parke's process.

313 (c) Chile salt petre ( $NaNO_3$ ) is the nitrate ore of sodium

314 (b) A water soluble complex of silver with a dilute aqueous solution of NaCN is sodium argentocyanide, in the cyanide process, the native from is crushed and treated with 0.1-0.2% solution of NaCN and aerated  
 $4Ag + 8NaCN + 2H_2O + O_2 \rightarrow 4Na[Ag(CN)_2] + 4NaOH$   
 Argentocyanide is soluble metal is recovered from the complex by reduction with zinc

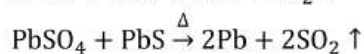
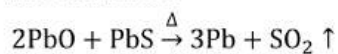
- 315 (c)  
Metals can not be extracted from all the minerals that is why all minerals are not ores
- 316 (a)  
Flux is used to fuse non-fusible impurities present in ore.
- 317 (d)  
 $\text{SnO}_2 + 2\text{C} \rightarrow \text{Sn} + 2\text{CO}$
- 318 (b)  
Wolframite is  $\text{FeWO}_4$ .
- 319 (c)  
 $\text{SiO}_2 + 2\text{C} \rightarrow \text{Si} + 2\text{CO} \uparrow$   
 $\text{Al}_2\text{O}_3 + 3\text{C} + \text{N}_2 \rightarrow 2\text{AlN} + 3\text{CO}$   
 $\text{AlN} + 3\text{H}_2\text{O} \rightarrow \text{Al(OH)}_3 + \text{NH}_3$
- 321 (b)  
Ge and Si both the elements are purified by Zone refining.
- 322 (b)  
It is a fact.
- 323 (d)  
It is a fact.
- 324 (c)  
It is a fact.
- 325 (d)  
Gypsum is  $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ .
- 326 (d)  
It is a fact.
- 327 (b)  
For purification of Ni in Mond's process.
- 328 (a)  
In electrolytic reduction, the oxides of highly electropositive metals are reduced at very high temperature
- 329 (d)  
CaO is hygroscopic agent.
- 330 (d)  
The slag obtained during the extraction of copper from copper pyrites is of  $\text{FeSiO}_3$ . It is carried out in smelting.  
 $\text{FeO} + \text{SiO}_2 \rightarrow \text{FeSiO}_3$   
(slag)
- 331 (a)  
Matte is a mixture of  $\text{Cu}_2\text{S}$  containing little FeS
- 332 (d)  
An ore is a mineral or aggregate of mineral from which a valuable constituent, especially a metal, can be profitably mined or extracted. All ores are minerals but all minerals are not ore.
- 333 (c)  
It is a fact.
- 334 (a)  
Pitch blende contains traces of radium.
- 335 (d)  
Roasting is mainly employed to remove volatile substances  
 $\text{S}_8 + 8\text{O}_2 \rightarrow 8\text{SO}_2 \uparrow$   
 $\text{P}_4 + 5\text{O}_2 \rightarrow \text{P}_4\text{O}_{10} \uparrow$   
 $4\text{As} + 3\text{O}_2 \rightarrow 2\text{As}_2\text{O}_3 \uparrow$
- 336 (d)  
Zn, Cd, Hg have low b.pt.
- 337 (a)  
Rutile is  $\text{TiO}_2$ .
- 338 (b)  
Argentite is  $\text{Ag}_2\text{S}$ .
- 339 (c)  
Alkali metals, alkaline earth metals and Al are extracted by electrolytic reduction.
- 340 (c)  
Wolframite is ferrous tungstate ( $\text{FeWO}_4$ ) which is magnetic in nature
- 341 (c)  
CaO is a basic flux.
- 342 (d)  
Cinnabar is an ore of Hg(HgS).
- 343 (c)  
Less reactive metals are found in native state (free state)
- 344 (c)  
Levigation (gravity separation) is based on the difference in the specific gravities of the gangue particles and the ore particles.
- 346 (d)  
Sodium has high reactivity towards water.
- 347 (b)  
Metals like, Na, K, Mg, Ca, Al etc are reduced by electrolytic reduction
- 348 (b)  
It is a fact.
- 349 (a)  
In the metallurgy of iron, when  $\text{CaCO}_3$  is added to blast furnace, it removes impurities from ore and forms slag.  
 $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$  (1070-1170 K)  
 $\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$  (1470 K)  
 $3\text{CaO} + \text{P}_2\text{O}_5 \rightarrow \text{Ca}_3(\text{PO}_4)_2$
- 350 (d)  
Mg alloys are lighter.
- 351 (c)  
It is definition of roasting.
- 352 (c)



Aluminium is mainly isolated from bauxite ( $\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$ ) ore which is generally contaminated with ferric oxide and silica

353 (d)

$\text{PbO}$  and  $\text{PbSO}_4$  get reduced by  $\text{PbS}$  itself which is already present in mixture, because the reduction takes place by mixture itself, hence is known as self reduction



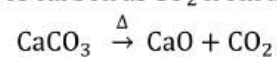
354 (b)

Calcination is a process in which the ore is heated strongly in the absence of air.

(i) It removes the volatile impurities like  $\text{CO}_2$ ,  $\text{SO}_2$ , organic matter, moisture from the ore.

(ii) It removes water from the hydrated ore.

(iii) It removes carbon as  $\text{CO}_2$  from a carbonate ore.



Lime stone

355 (c)

It is a fact.

356 (d)

The temperature of the slag zone in the metallurgy of iron using blast furnace is 800-1000°C

357 (d)

A natural crystalline form of blue, transparent corundum ( $\text{Al}_2\text{O}_3$ ). The colour being due to traces of cobalt and other metals.

